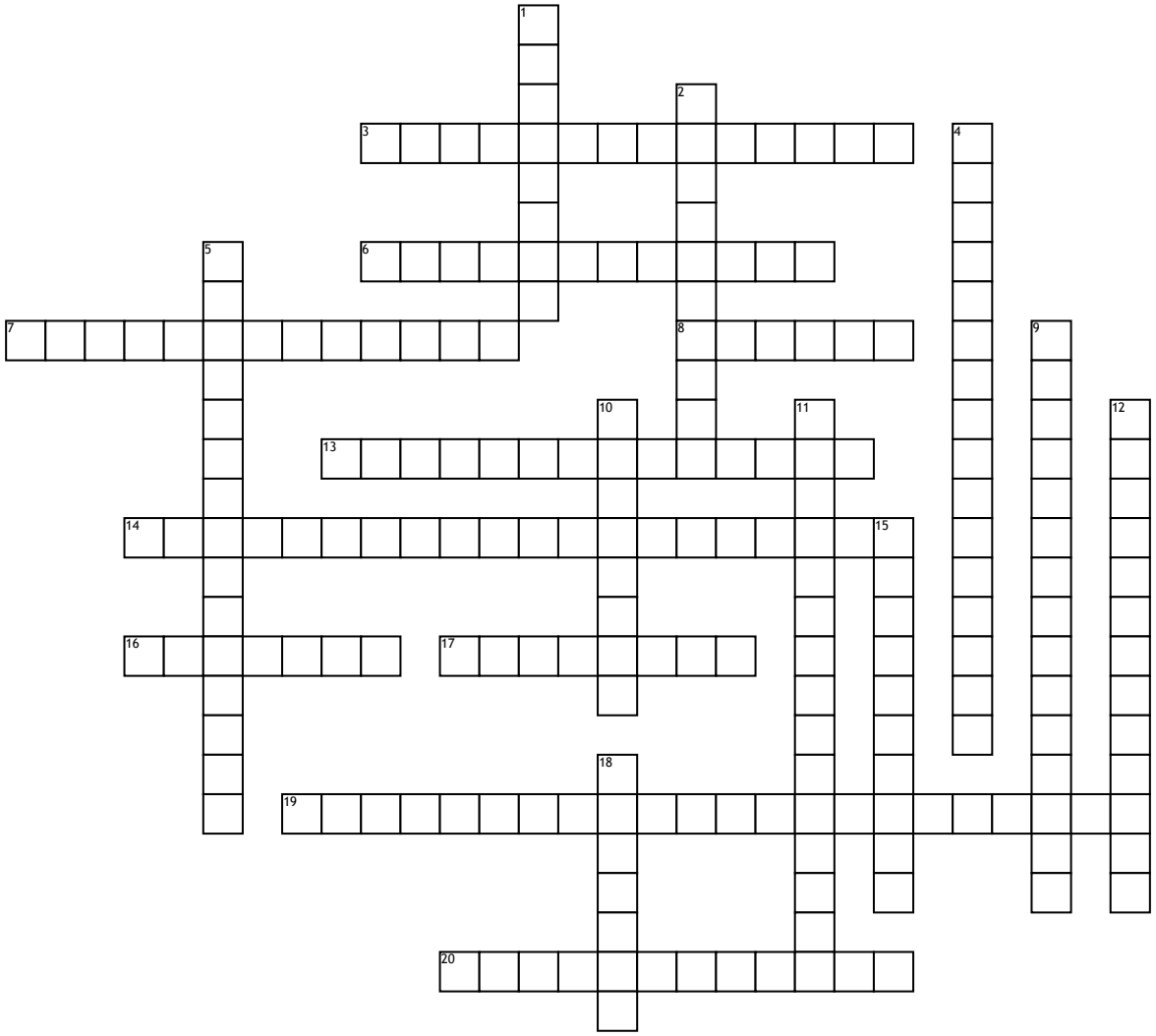
Name: _____ Date: _____ Period: _____

chemistry



Across

3. The inward force or pull that tends to minimize the surface area of the liquid; Water tends to drop in spheres; Water not attracted to air molecules

6. Conduct electricity when dissolved in a solution Ionic Compounds

7. mass solute x 100 / mL solution

8. The substance that is being dissolved

13. As water cools, it starts to contract, causing mass to decrease while volume remains constant.; D = M/V; 10/10 = 1 5/10 = ½; At 0°C ice has 10% greater volume than water, therefore a lower density.

14. The greater the intermolecular forces the more spherical the drop

16. The medium that is dissolving another substance

17. Homogenous mixture; solid, liquid, or gas

19. Extensive bonding requires large amounts of heat to disrupt; Boiling point: 100°C

20. The amount of heat required to change the temperature by 1°C; High Specific Heat = 4.18 J/g °C; Helps to moderate air temperatures

Down

1. m = mol solute / kg solvent

2. S1 /P1 = S2 / P2 The solubility of a gas in a liquid is directly proportional to the partial pressure of the gas on the surface of the liquid

4. Pressure created as molecules that escape from the liquid phase to the gaseous phase; Hydrogen Bonding ; Holds molecules tightly, does not allow escape

5. Do not conduct electricity Covalent Compounds

9. Creates a large network between adjacent molecules; Forms a 'skin'

10. M = mol solute / L solution

11. vol solute x 100 / vol solution

12. The FIZZING that happens when you open the pop; the release of gas from a solution

15. Particles > 100 nm; settle to the bottom (sand & water)

18. Particles 1-100 nm; suspended throughout the medium Scattering of light