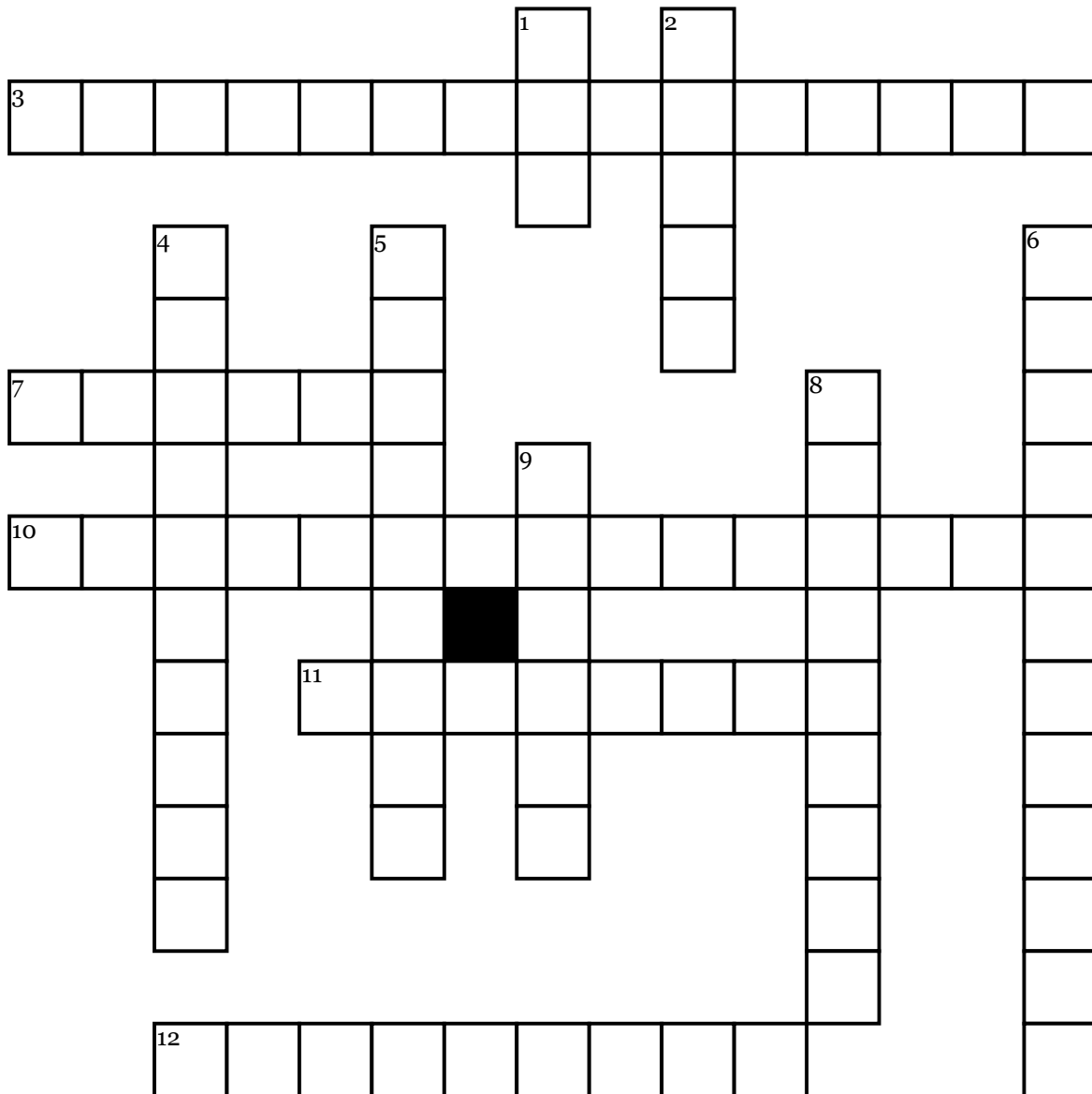


bonding elements



Across

- 3.** The total number of electrons that an atom either gains or loses in order to form a chemical bond with another atom.
- 7.** Atoms that lose electrons and have a positive (+) charge.
- 10.** The electrons found in the outermost shell of an atom. This is where the chemical reactions and bonding takes place.
- 11.** Bond This is the type of bond where electrons are shared between atoms.

- 12.** These are found to the right of the staircase and form covalent bonds with each other, or take electrons away from metals to form ionic bonds and become negative (-) charged.

Down

- 1.** Atoms with either a positive (+) or negative (-) electrical charge.
- 2.** Atoms that gain electrons and have a negative (-) charge.
- 4.** These touch the staircase in the periodic table and have properties of both metals and non metals.

- 5.** Type of bond where atoms give electrons to other atoms and there is an attraction between oppositely-charged ions.
- 6.** The outermost shell of an atom.
- 8.** Atoms want to have eight (8) electrons in their outer shell to become stable.
- 9.** These are found to the left of the staircase and form ionic bonds with non-metals. They give away electrons and become positive (+) charged. Except for Hydrogen.