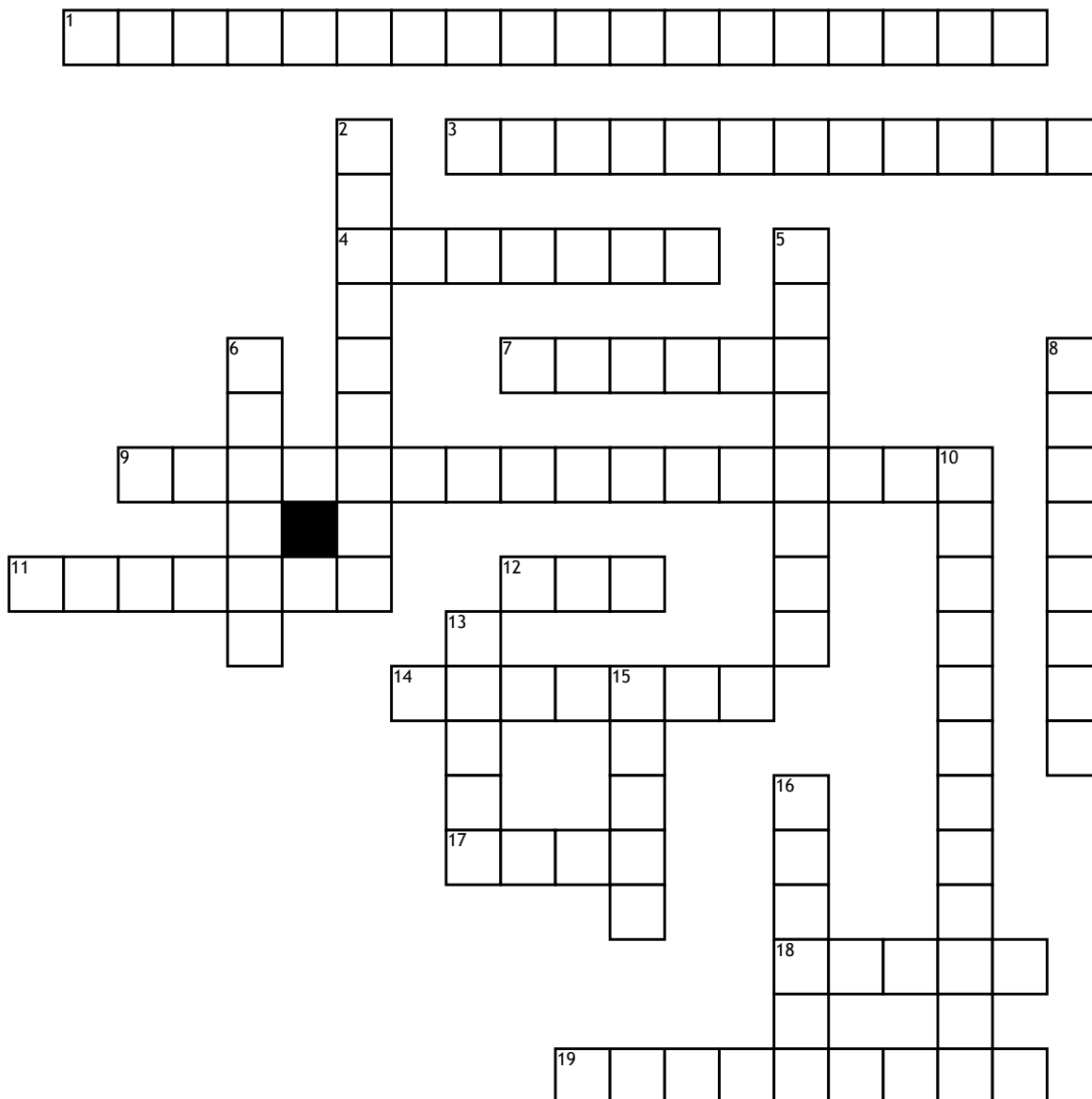


# Year 9 revision - Units 4 and 5



**Across**

- 1. Mendeleev left gaps in his Periodic Table because he put the elements in groups with similar what? (8,10)
- 3. What did Mendeleev change the order of? (6,6)
- 4. What is the mnemonic for the diatomic molecules? (7)
- 7. The vertical columns of elements are called what? (6)
- 9. Which group of elements act as catalysts, produce coloured compounds and have ions of different charges? (10,6)
- 11. What is the charge on the Fe (II) ion? (3,4)

- 12. What is the formula of sodium fluoride? (3)
  - 14. The horizontal rows of elements are called what? (7)
  - 17. Do metals lose or gain electrons to get a full outer shell? (4)
  - 18. What type of bond forms between a metal and non-metal? (5)
  - 19. What happens to the reactivity as you go down group 1? (9)
- Down**
- 2. What happens to the reactivity as you go down the Halogens? (9)
  - 5. Metals form ions with which charge? (8)

- 6. Atoms become ..... and unreactive when they form ions with a full outer shell (6)
- 8. When group 1 metals react with water, which gas is produced? (8)
- 10. What is the name of a compound formed between sodium and iodine? (6,6)
- 13. Two negative ions will do what? (5)
- 15. The group number is the same as the number of electrons in which shell of an atom? (5)
- 16. What is the ion of a halogen called? (6)