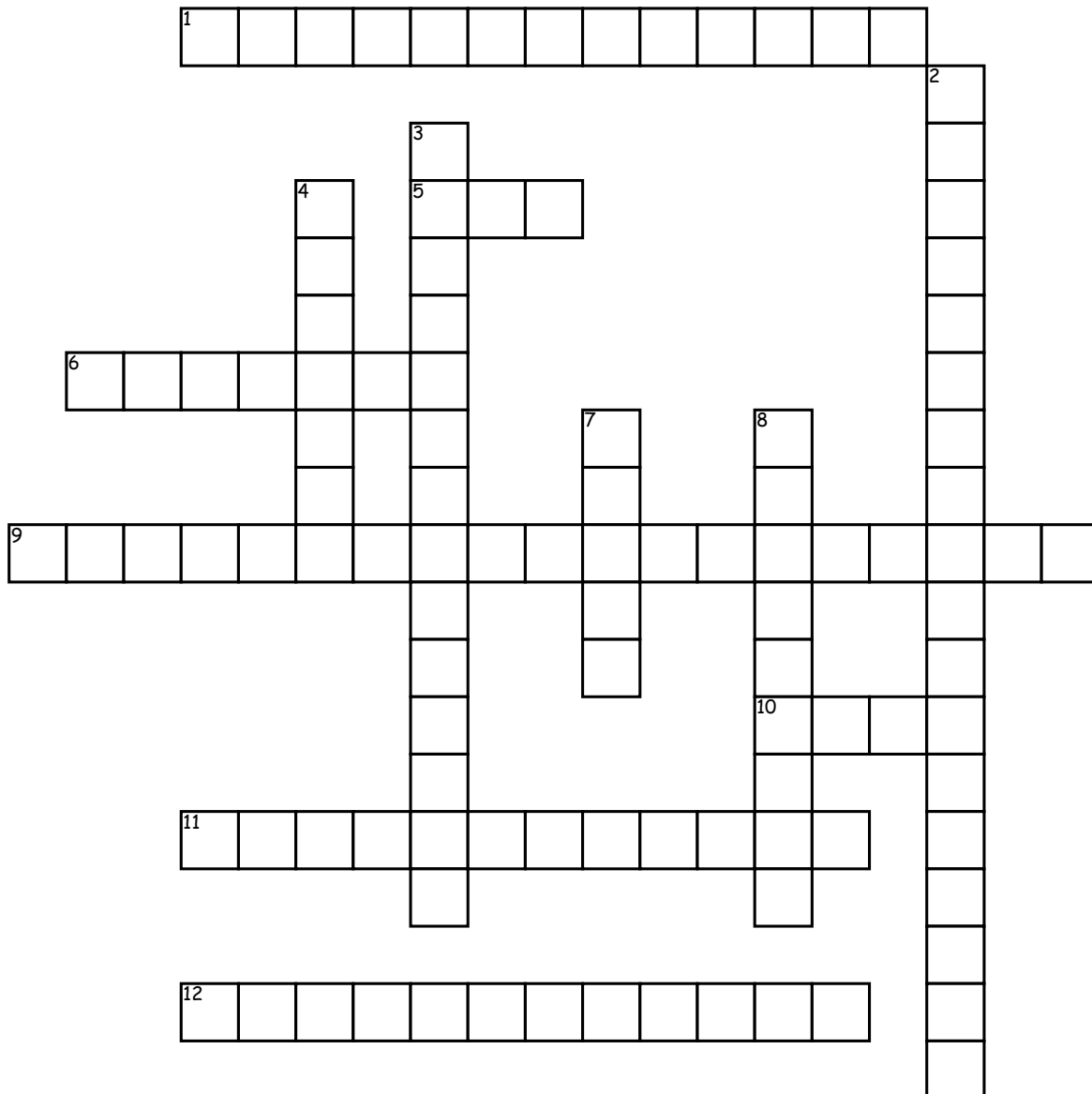


Unit 5 Bonds



Across

1. A layer of electrons surrounding the nucleus of an atom.

5. A charged particle formed when an atom gains or loses one or more electrons. In an ion, the number of electrons is different from the number of protons in the nucleus

6. A piece of a solid element or compound with a lattice structure. The shape of the crystal is due to the shape of the lattice.

9. A bond between two atoms which involves sharing two pairs of electrons.

10. A bond is what holds ions or atoms together. The forces that make the bond are different in different examples.

11. A bond between two atoms which involves sharing a pair of electrons.

12. An electron which is not involved in a chemical bond, and can move when there is a charge.

Down

2. We can use dots (I) and crosses (N) to represent electrons in a covalent bond.

3. A molecule with covalent bonds. It contains a small number of atoms.

4. A regular repeating pattern of atoms which makes a solid element or compound.

7. A lattice structure which has many atoms with covalent bonds

8. A bond between two ions. The ions have opposite charges.