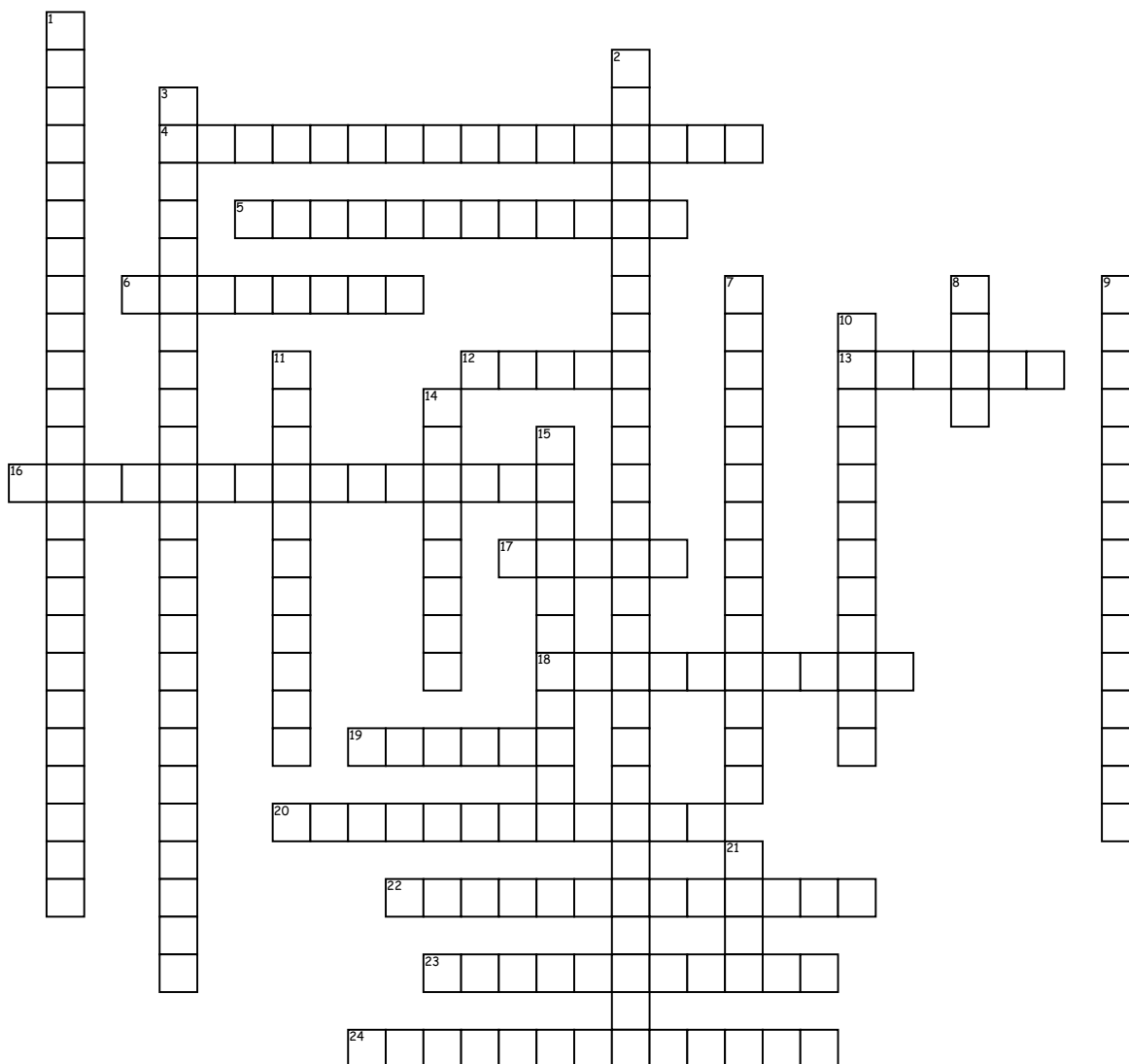


Thermochemistry



Across

4. "heat" of the reaction

5. everything else

6. In going from a particular set of reactants to a particular set of products the change in enthalpy is the same whether the reaction takes place in one step or in a series of steps

12. any kind of push or pull exerted on an object

13. the capacity to do work or to produce heat

16. the change in enthalpy that accompanies the formation of 1 mol of the compound from its elements with all of the substances in their standard states

17. Si unit for energy

18. reaction which releases energy; energy flows out of system

19. the part of the universe under examination in a chemical reaction

20. heat capacity of one gram of a substance

22. energy of motion

23. measure of heat flow

24. A function/property of a system that depends only on its present state (conditions). It is independent of the pathway (how the system arrived at that state)

Down

1. Energy can be converted from one form to another but cannot be created or destroyed

2. difference between the enthalpy of the productions and the reactants

3. 9.8 m/s^2

7. sum of all the kinetic energy and potential energy of all its components

8. the energy used to cause an object with mass to move against a force

9. energy due to position or chemical composition

10. temperature change experienced by an object when it absorbs a certain amount of heat

11. device used to measure heat flow

14. equal to energy (E) + pressure(P)volum(V)

15. reaction which absorbs energy; energy flows into system

21. the energy used to cause the temperature of an object to increase