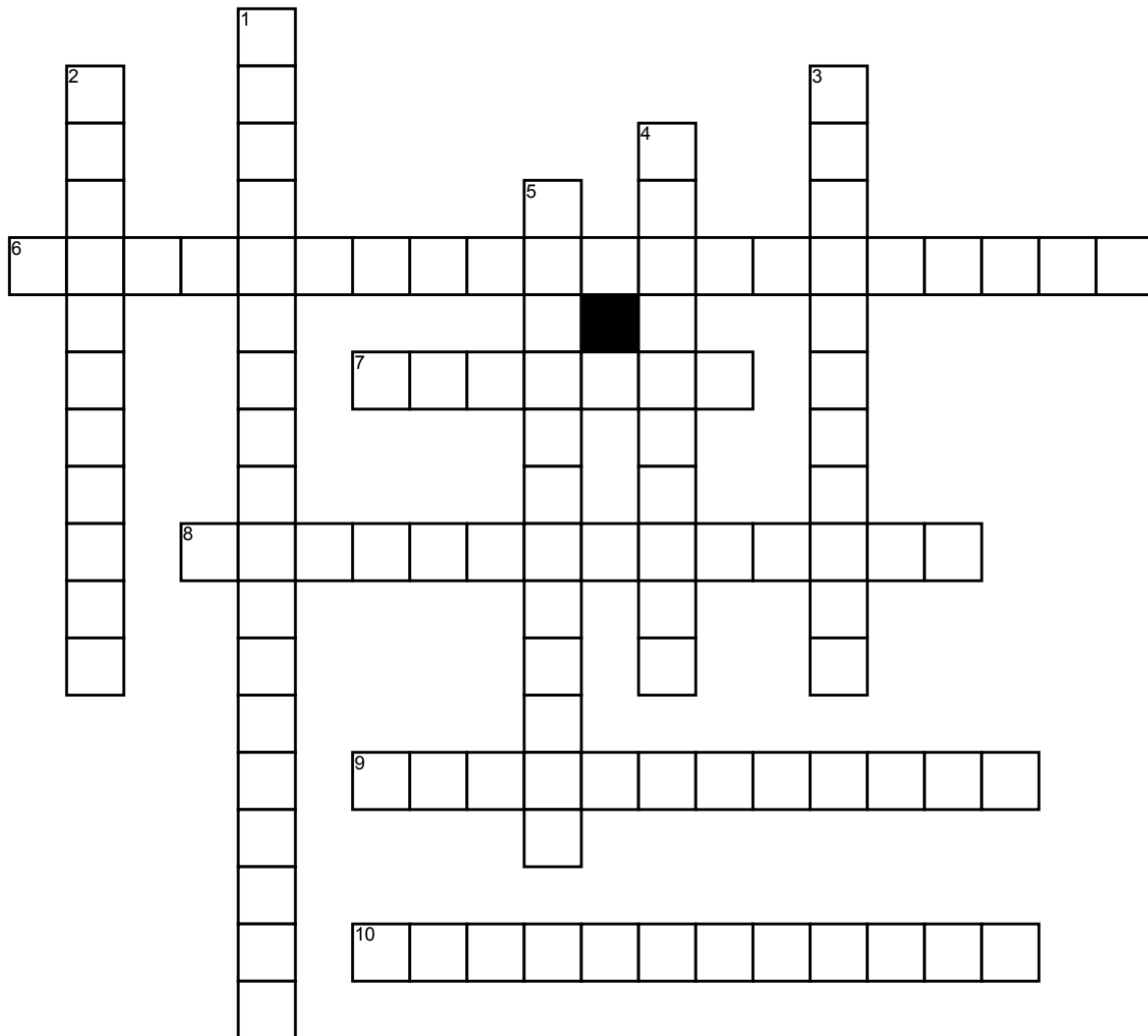


Thermal capacity and Melting & Boiling



Across

- 6.** $c = Q/(mT)$
- 7.** The process in which the heat energy absorbed changes a substance from a solid into a liquid. The substance itself does not experience a temperature change; all thermal energy added to system is used as latent heat to overcome the forces of attraction holding the particles in their solid lattice.
- 8.** A measure of the tendency of a liquid to turn into a gas.
- 9.** The temperature at which the forces of attraction between particles are overcome.
- 10.** The temperature at which the vapour pressure is equal to the atmospheric pressure.

Down

- 1.** The quantity of heat energy absorbed or released when 1kg of a substance changes state without changing temperature
- 2.** The process in which water turns into a gas without boiling.
- 3.** Vapour pressure increases with increasing
- 4.** The quantity of heat energy absorbed or released when a substance changes state without changing its temperature.
- 5.** The heat required to raise an object's temperature by one Kelvin