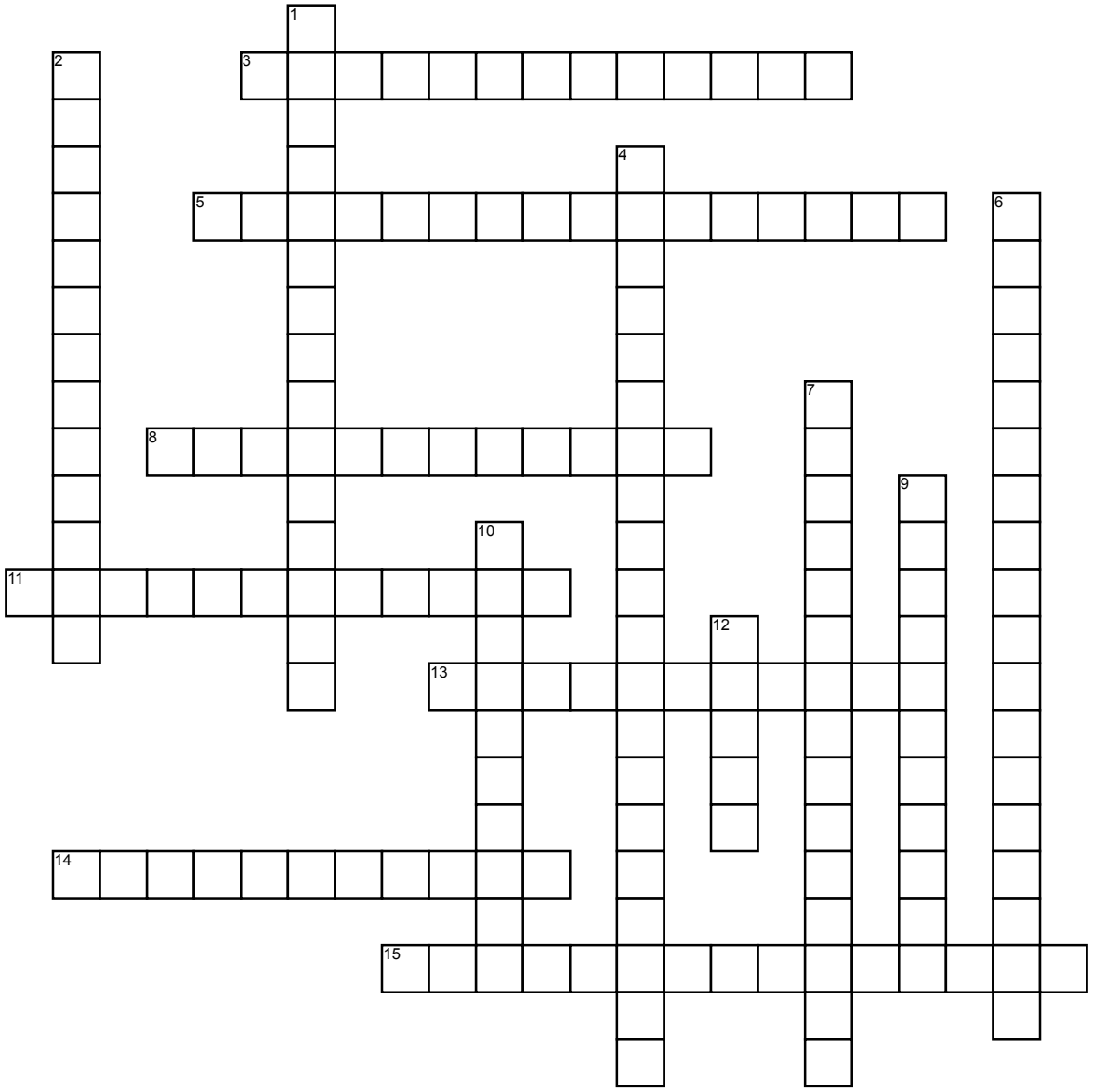
The Chemistry of Acids & Bases



Across

3. The substance that remains after an acid has released the hydrogen ion.

5. Substances that dissociate or ionize only slightly when added to water

8. Substance that when dissolved in water form ions, does conduct an electric current

11. The process by which the ions of salts separate from each other when dissolved in water.

13. Acid that contains hydrogen and one other element (non-metal)

14. A substance that can act as either an acid or a base

15. Atoms or molecules that do not form ions when dissolved in water, do not conduct electric current

Down

1. Acids that contain hydrogen and 2 other elements (non-metals)

2. The substance that remains after the base has accepted the hydrogen ion

4. Defines acid as a proton donor and a base as a proton acceptor

6. Substances that dissociate or ionize almost 100% when added to water

7. Defines an acid as a substance that when dissolved in water produces H+1 ions and a base, dissolved in water creates OH-1 ions

9. Defines acid as an electron pair acceptor and base as an electron pair donor

10. The formation of ions when neutral molecules like acids and bases react with water

12. Chemical sthat do not release H+1 ions or OH-1 ions into the solution when they dissociate