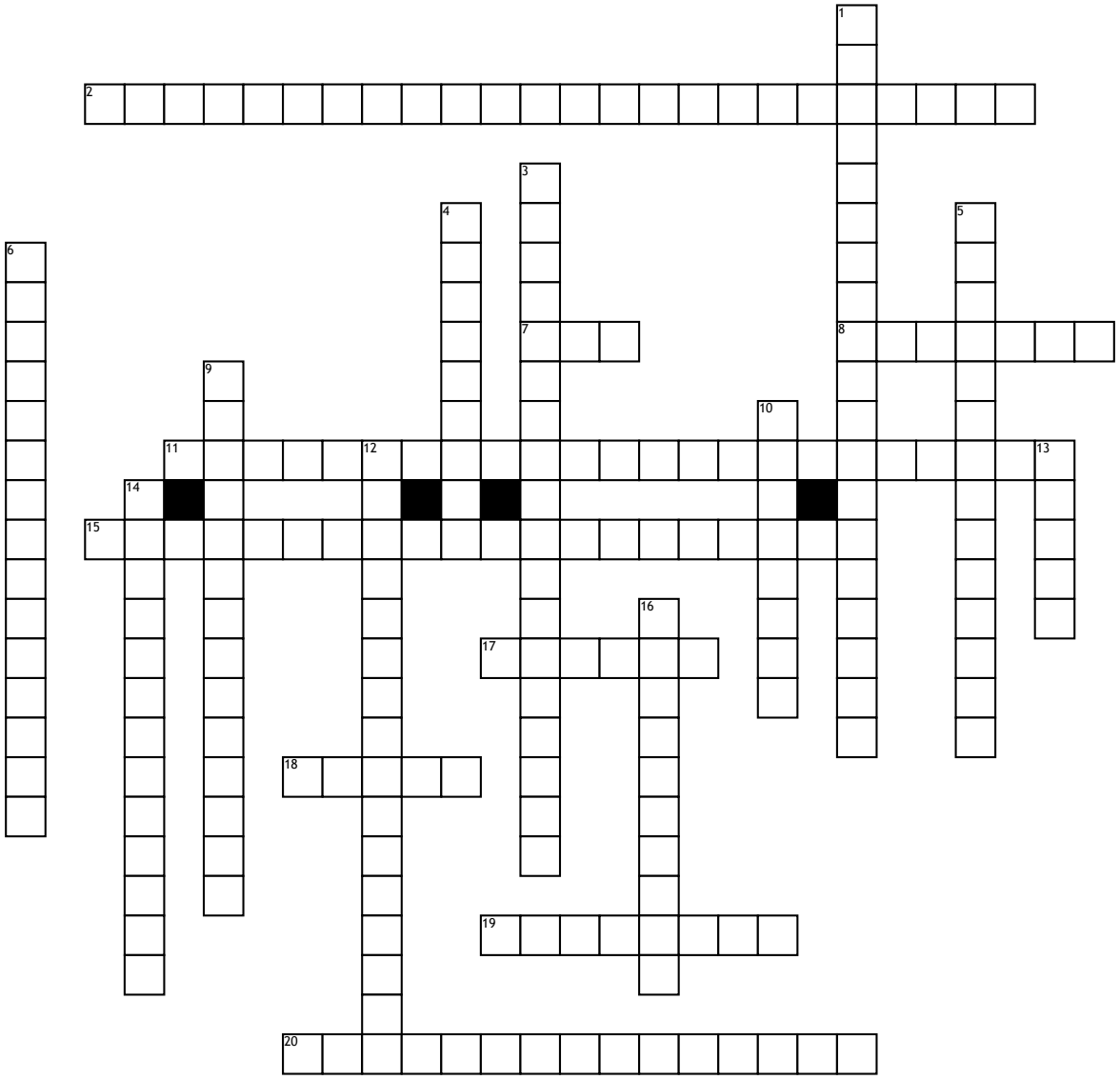
Name: _____ Date: _____

States of Matter



Across

2. States that, regardless of the amount, a compound of the same elements in the same proportion by mass

7. A form of matter that flows to conform to the shape of its containers entire volume, and is easily compressed

8. A pure substance that cannot be broken down into simpler substances by physical or chemical means

11. States that mass is neither created nor destroyed during a chemical reaction but is conserved

15. One that does not have a uniform composition and in which the individual substance remain distinct

17. A form of matter that flows, has constant volume, and takes the shape of its container

18. Gaseous state of a substance that is a liquid or a solid at room temperature

19. A chemical combination of two or more different elements; can be broken down into simpler substance by chemical means and has properties different from those of its component elements

20. A type of change that alters the physical properties of a substance but does not change its composition

Down

1. A physical property that remains the same no matter how much of a substance is present

3. One that has a uniform composition throughout and always has a single phase; also called a solution

4. A form of matter that has a uniform and unchanging composition; also known as a pure substance

5. A process involving one or more substances changing into new substances; also called a chemical reaction

6. A separation technique that produces pure solid particles of a substance from a solution that contains the dissolved substance

9. The physical forms, in which all matter naturally exists on Earth most commonly as a solid, a liquid, or a gas

10. A uniform mixture that may contain solids, liquids, or gases; also called a homogeneous mixture

12. The ability or inability of a substance to combine with or change into one or more new substances

13. A form of matter that has its own definite shape and volume, is incompressible, and expands only slightly when heated

14. A percentage determined by the ratio of the mass of each element to the mass of the compound

16. A technique that used a porous barrier to separate a solid from a liquid