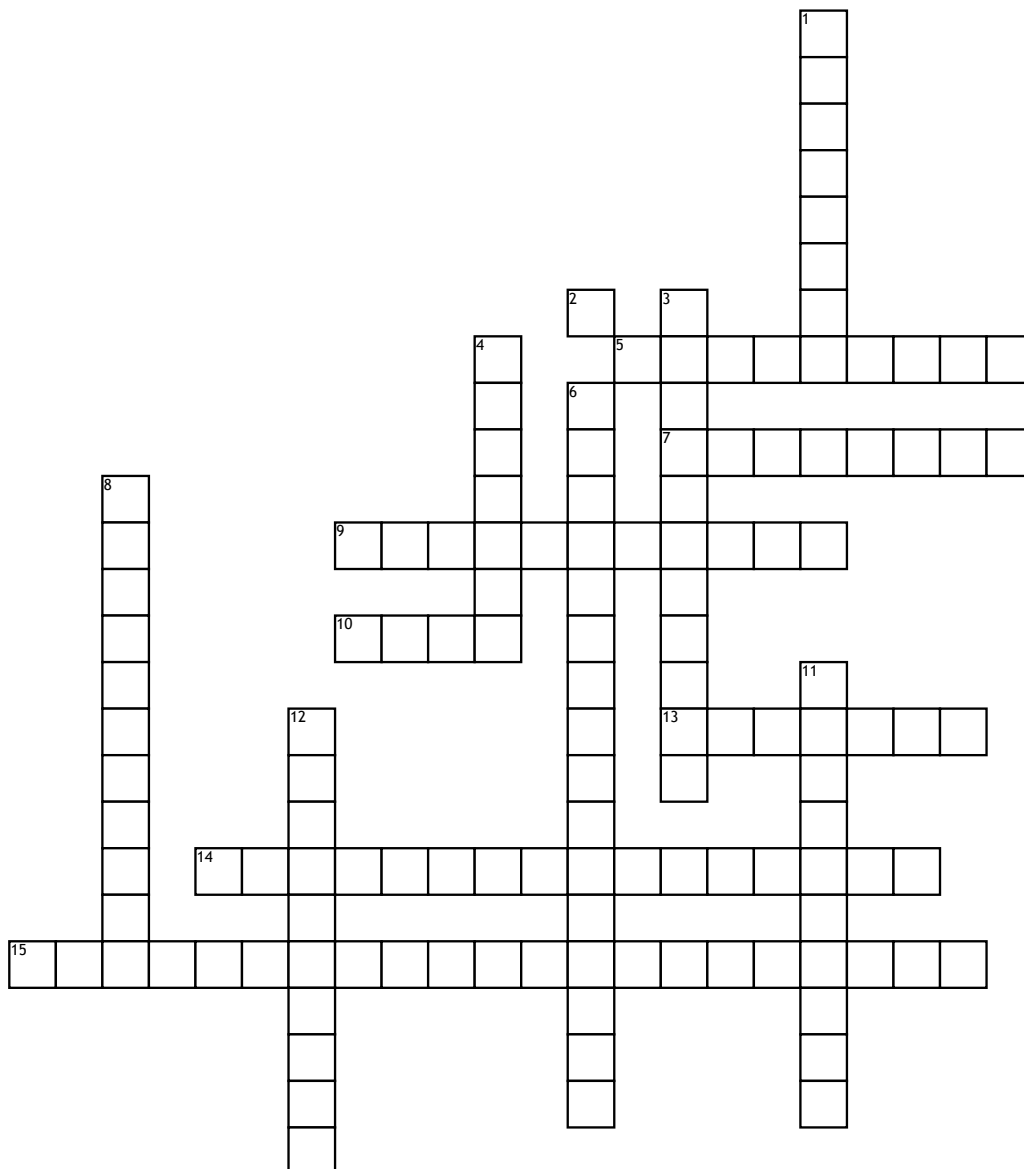


Rate and Equilibrium



Across

5. When the K_{eq} is less than 1, that means more _____

7. When the K_{eq} is more than 1, that means more _____

9. K_{eq} is the _____ constant.

10. _____ Substances are not included in the final equilibrium equation (liquids and solids)

13. This expresses the relationship between the rate of reaction and the concentration of the reaction.

14. The minimum amount of energy required by resting particles in order to form the activated complex.

15. If a dynamic equilibrium is disturbed by the changing of conditions the position of equilibrium moves to relieve the change is called _____

Down

1. _____ Increase the speed of reactions without being permanently altered by changing the reaction mechanism so less activation energy is needed.

2. What letter represents the rate constant?

3. The value of K_{eq} is constant only at a specified _____

4. Collision Theory says atoms, molecules, must _____ in order to react.

6. Also known as a transition state

8. The type of reaction that absorbs heat.

11. What type of reaction can occur in both forward and reverse directions?

12. The Type of reaction that releases heat