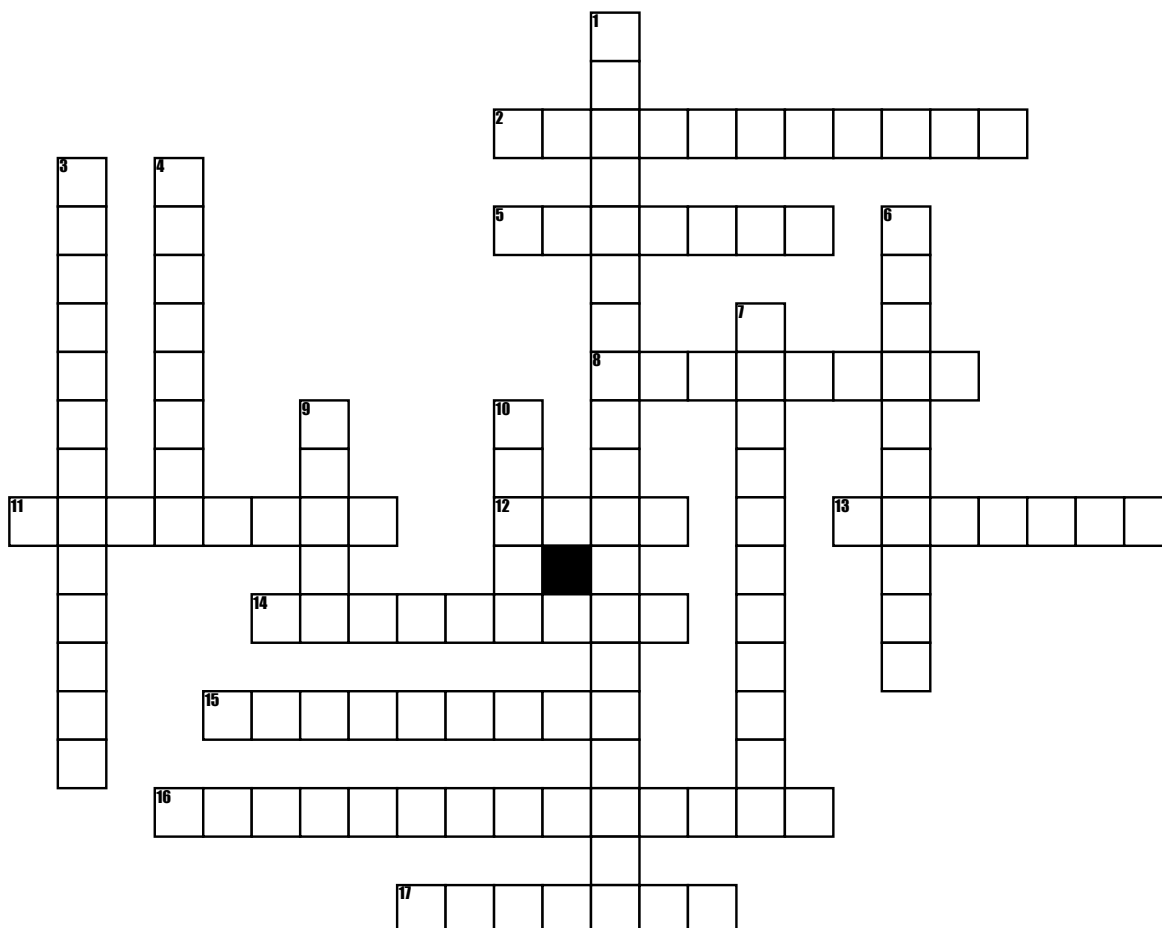


Rate and Equilibrium



Across

2. Chemical Reactions often reach a _____ point or equilibrium?

5. Chemical equilibrium is a state in which forward and reverse reactions _____ each other?

8. What happens to the heat of a Exothermic reaction?

11. $K_{eq} > 1$: is more _____?

12. Reaction Rate is the change in concentration of a reactant or product per unit of _____?

13. How much energy is required for Activation Energy?

14. What is the theory in which atoms, ions, and molecules must collide in order to react?

15. $K_{eq} < 1$: is more _____?

16. What are not included in the final equilibrium equation?

17. Increasing the temperature of a reaction increases _____ energy?

Down

1. K_{eq} is the _____?

3. A rate of law express (revel) the relationship between the rate of chemical reaction and the _____ of the reaction?

4. What happens to the heat of a Endothermic reaction?

6. A _____ reaction is a chemical reaction that can occur in both the forward and reverse directions such as the formation of ammonia.

7. The value of K_{eq} is constant only at a specified _____?

9. Equilibrium constant is the numerical value of the _____ of product concentrations to reactant concentrations?

10. Chemical equilibrium take place at equal _____?