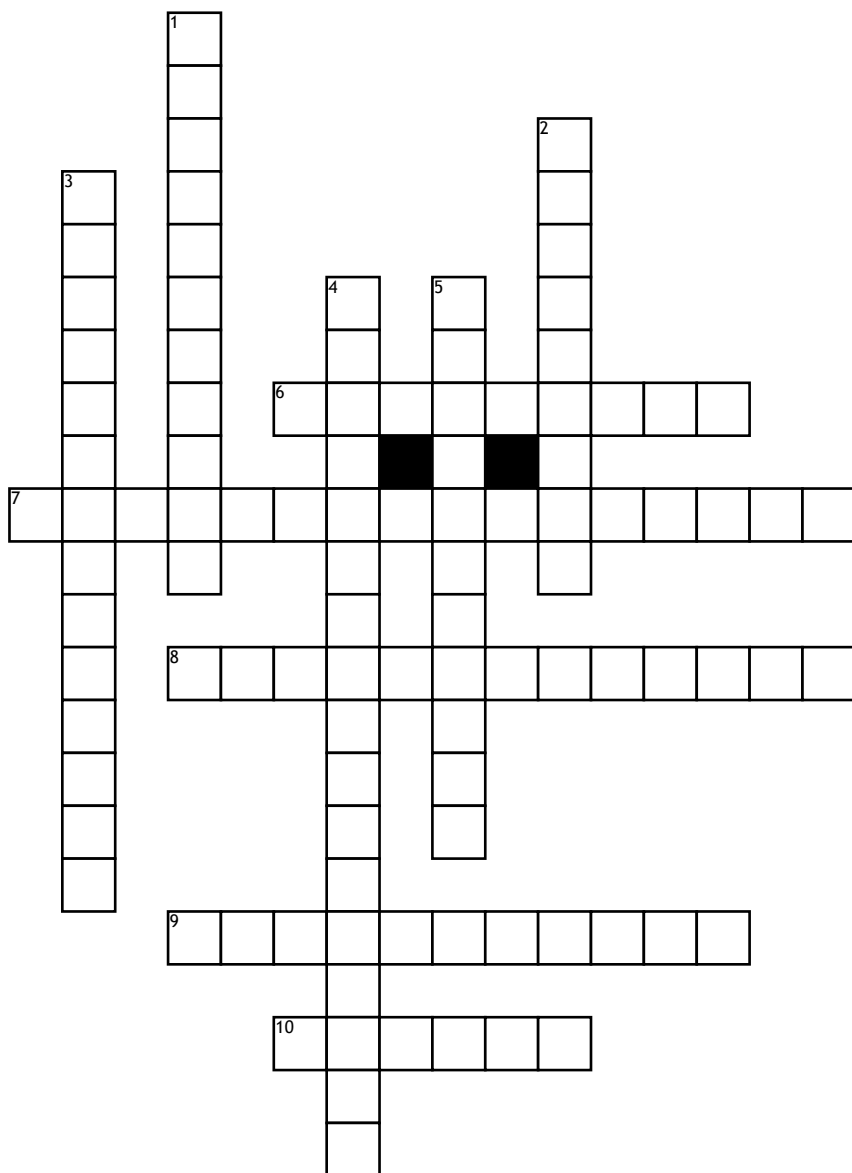


# Periodicity and Ionisation Energy



## Across

6. The atomic radii \_\_\_\_\_ across a period
7. The energy required to remove 1 mole of electrons from 1 mole of gaseous atoms.
8. Carbon and silicon have a \_\_\_\_\_ structure with strong covalent bonds and so have a high boiling point.
9. The attraction of electrons towards the nucleus
10. The second IE is \_\_\_\_\_ than the first IE.

## Down

1. The total distance from an atom's nucleus to the outermost orbital of electron
2. The atomic radii \_\_\_\_\_ down a group
3. Simple molecules have weak \_\_\_\_\_ forces of attraction and so have a low boiling point.
4. The repulsion of outer shell of electrons by inner shells of electrons.
5. The trends in the properties of elements across a period