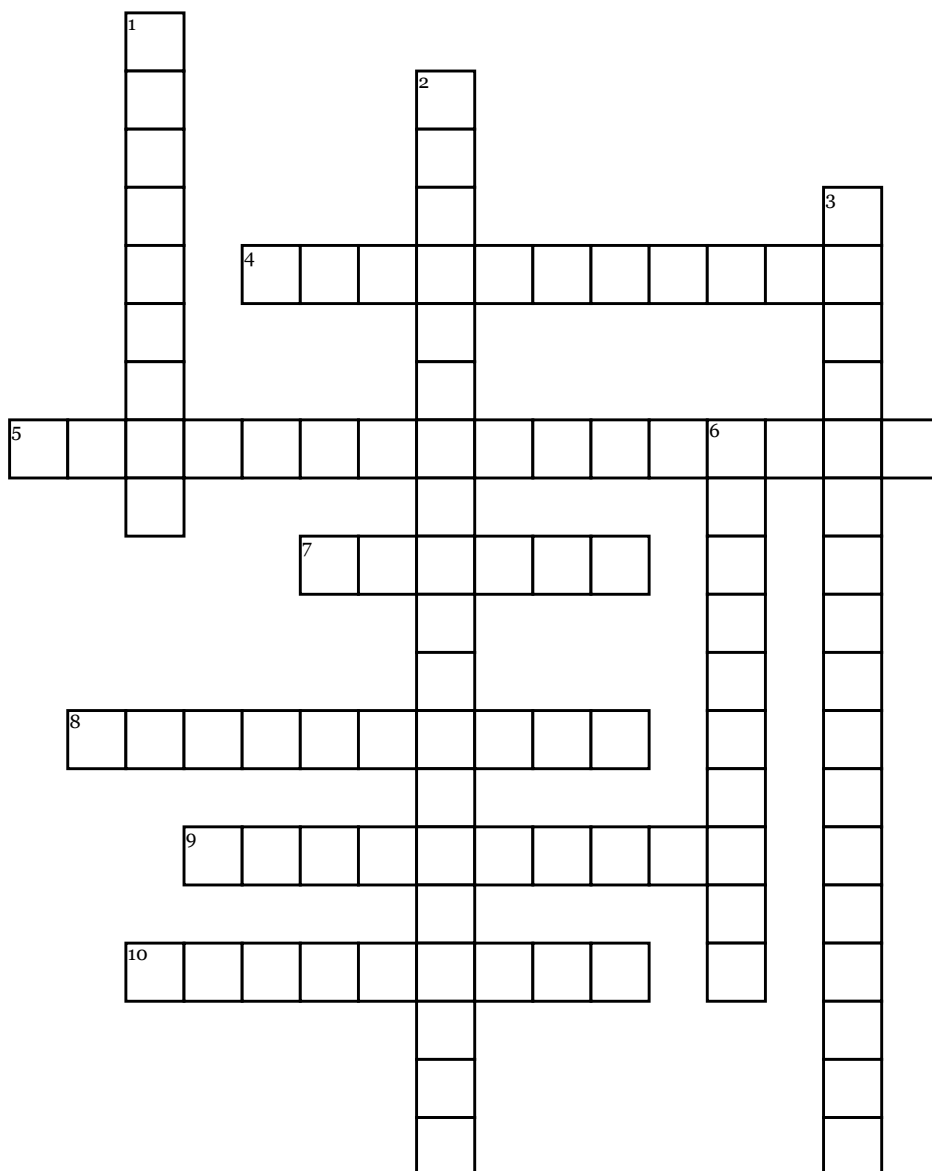


Name: _____ Date: _____ Period: _____

Periodic Table Crossword Puzzle



Across

4. any of the elements lithium, sodium, potassium, rubidium, cesium, and francium, occupying Group IA (1) of the periodic table. They are very reactive, electropositive, monovalent metals forming strongly alkaline hydroxides.

5. is of an atom or molecule is defined as the amount of energy released or spent when an electron is added to a neutral atom or molecule in the gaseous state to form a negative ion.

7. arising from the arrangement of the periodic table, provide chemists with an invaluable tool to quickly predict an element's properties.

8. is the process by which an atom or a molecule acquires a negative or positive charge by gaining or losing electrons to form ions, often in conjunction with other chemical changes.

9. a chemical element (as boron, carbon, or nitrogen) that lacks the characteristics of a metal and that is able to form anions, acidic oxides, acids, and stable compounds with hydrogen.

10. is an element that doesn't have the characteristics of metal including: ability to conduct heat or electricity, luster, or flexibility. An example of a nonmetal element is carbon.

Down

1. introduced by Niels Bohr and Ernest Rutherford in 1913, depicts the atom as a small, positively charged nucleus surrounded by electrons

2. any of the six highly reactive chemical elements in group IIA of the periodic table including, in order of increasing reactivity, beryllium, magnesium, calcium, strontium, barium, and radium: an oxide of any of these metals is called

3. is a measure of the tendency of an atom to attract a bonding pair of electrons.

6. any of the gaseous elements helium, neon, argon, krypton, xenon, and radon, occupying Group 0 (18) of the periodic table. They were long believed to be totally unreactive but compounds of xenon, krypton, and radon are now known.