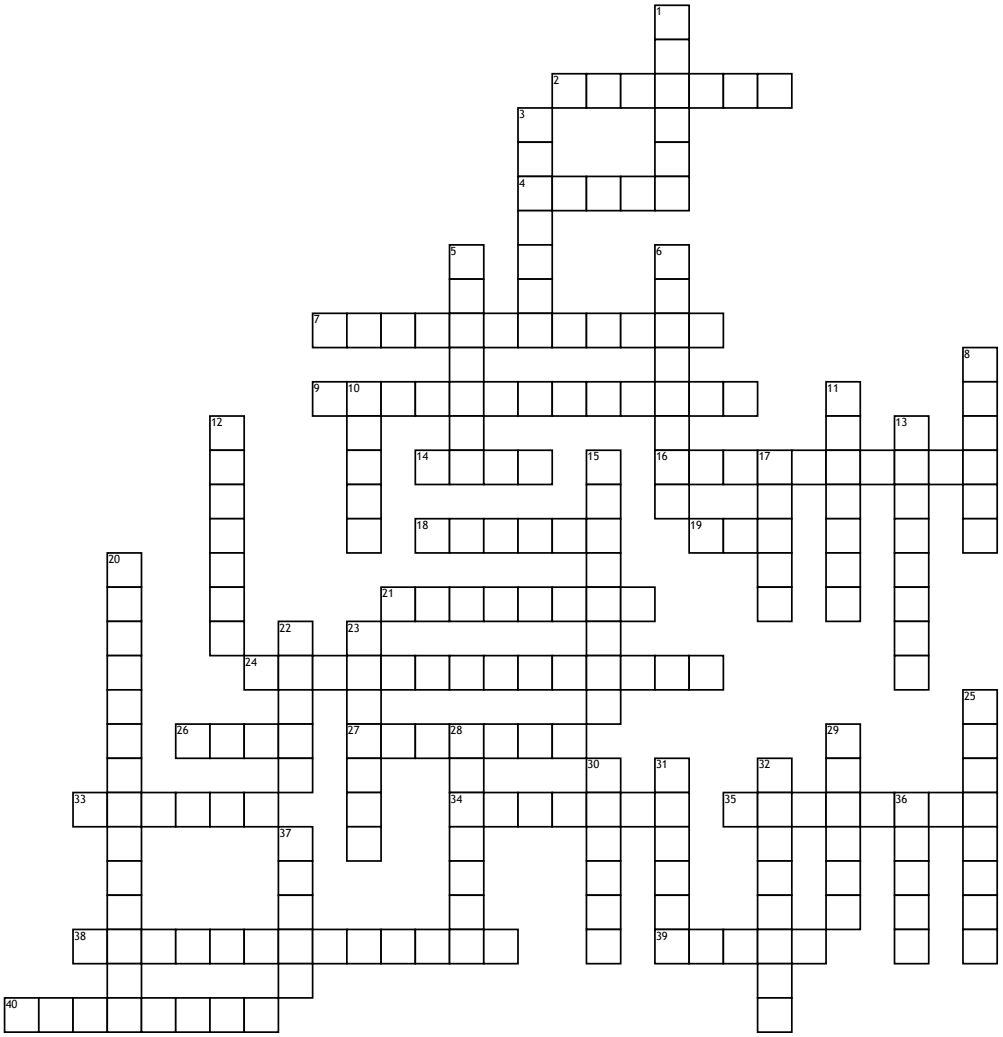
Name: _____ Date: _____

Our Atomic World LP1



Across

2. Nuclear reaction producing lighter elements and energy.

4. Reaction that continues to feed itself and keep itself going is a _____reaction.

7. Becquerel's discovery was caused by a piece of Uranium ore being put in his desk with a_____Plate.

9. The process by which a substance gives off penetrating particles.

14. A negative particle identical to an electron but radiating from a decaying nucleus is a_____particle.

16. The total number of protons plus neutrons in an atomic nucleus also known as atomic weight.

18. State of matter where the atoms are able to move around and take the shape of the container.

19. State of matter free to fill in all the space of it's container.

21. A tiny particle with a negative electrical charge which orbits the nucleus of an atom.

24. Professor who accidentally discovered radioactivity

26. The smallest particle of matter still retaining the characteristics of the element

27. The neutral or no charge particle of an atom

33. A positive charged particle in the nucleus of an atom

34. The class of matter in which there are 118 different varieties on the Periodic Table

35. A nuclear particle of the mass of an electron but with a possitive charge.

38. List of elements and their properties, symbol, name and atomic number

39. A basic unit of measuring the amount of radiation being given in one second is named the_____after Marie Curie.

40. A neutral particle of nearly zero mass, produced by decaying nuclei.

Down

1. nuclear reaction producing heavier elements and energy.

3. Energy produced by means of a nuclear reactor is called_____energy.

5. A form of an atom that has the same atomic number and same chemical properties but different atomic mass.

6. A spontaneous nuclear reaction like a bomb is called _____ mass explosion.

8. One way of measuring amounts of radiation is the _____ Cloud Chamber.

10. A radiation particle coming from the nucleus of some atoms is an _____ particle

11. The center of an atom

12. What opposite poles of a magnet do.

13. A bundle of energy that is released from nuclei and can pass right through a person.

15. First element of the periodic table. Symbol H

17. A particle one-eighth the mass of protons. there charge can be positive +, negative - or zero 0.

20. Substances which, after being exposed to sunlight, will glow in the dark.

22. Physicist who won the Nobel Prize in 1938 and helped create the first nuclear reactor.

23. Radioactive element number 92 on the Periodic Table with a symbol of U.

25. The other unit of measuring radiation is called the_____ after William Roentgen.

28. The effect heated water will have on the environment when returned to it's source is called_____effects.

29. Instrument used to determine the intensity of a radioactive sample is called a_____counter.

30. Second element on the periodic table. Symbol He

31. The number of protons in an atom is also known as the elements_____ Number.

32. Marie and Pierre Curie discovered a new radioactive element and named it after her homeland. Element 84, symbol Po.

36. What like poles of a magnet do.

37. State of matter that has a definite shape,size and mass

Word Bank

Fermi	atom	fission	nucleus	Attract	photographic	meson
proton	chain	curie	solid	Geiger	positron	thermal
atomic mass	Helium	Liquid	beta	neutrino	gas	element
Wilson	radioactivity	fusion	isotope	nuclear	roentgen	critical
alpha	atomic	electron	phosphorescent	neutron	Periodic Table	uranium
repel	Henri Becquerel	Polonium	gamma ray	Hydrogen		