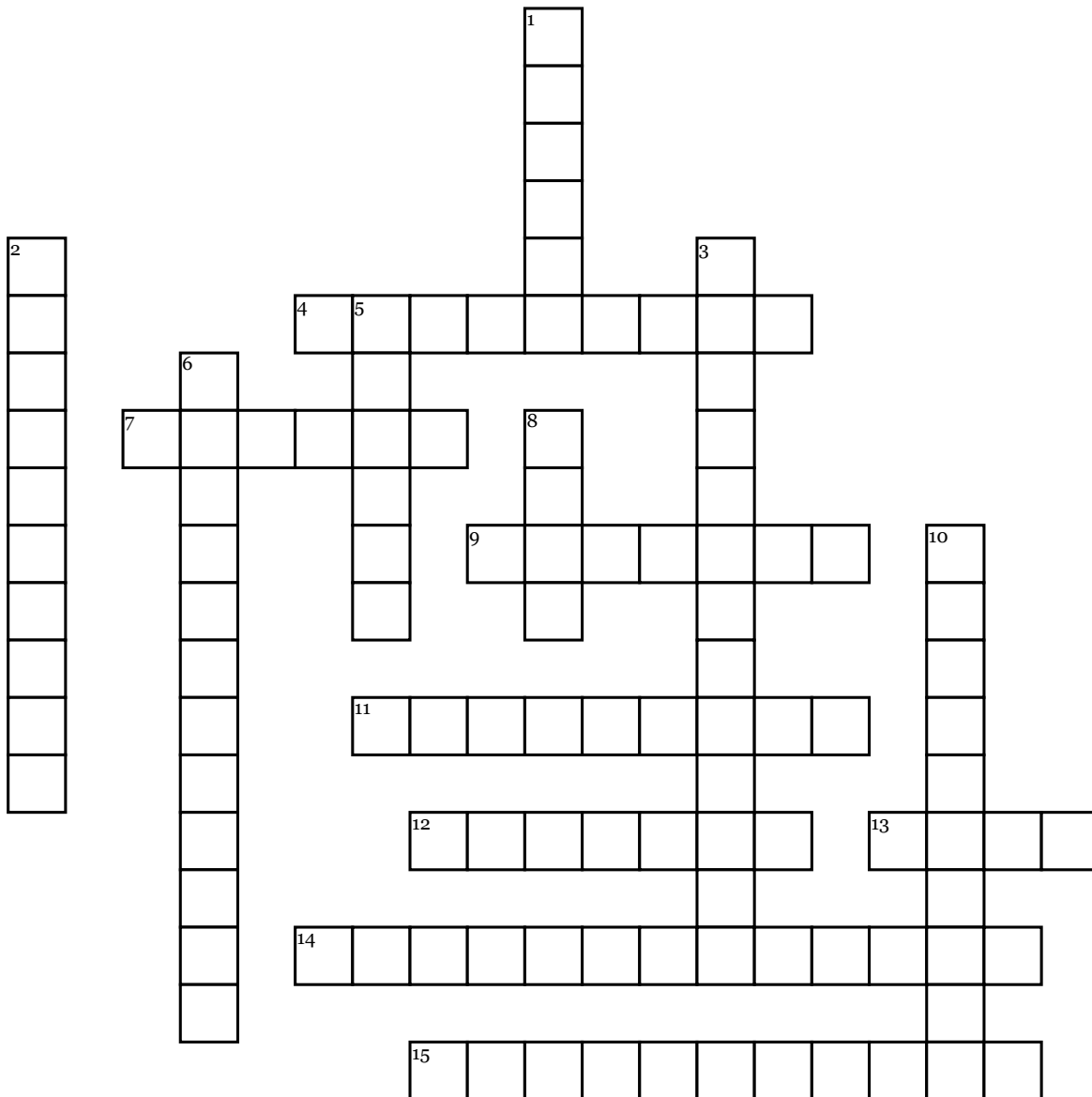


Metal Structures



Across

4. Transition metals have high _____

7. Most of the commonly used metals are there, including iron, _____, silver and gold.

9. This means that the particles are close together and in a _____ arrangement.

11. All metals are _____, meaning they can be hammered into shape.

12. That is why they have high _____ points and boiling points.

13. Transition metals have _____ melting points - but mercury is a liquid at room temperature

14. There are strong _____ forces holding the particles together.

15. Solid metals are _____.

Down

1. They are less reactive than _____ metals such as sodium

2. Transition metals are good _____ of heat and electricity

3. The particles in a metal are held together by strong _____

5. It takes a lot of _____ to separate the particles.

6. Metals have loose electrons in the outer shells which form a 'sea' of detached negative charge around the close-packed _____.

8. The advantage of malleability is that metals have many _____, eg for making jewellery.

10. The elements in the centre of the periodic table, between groups 2 and 3, are called the _____ metals.