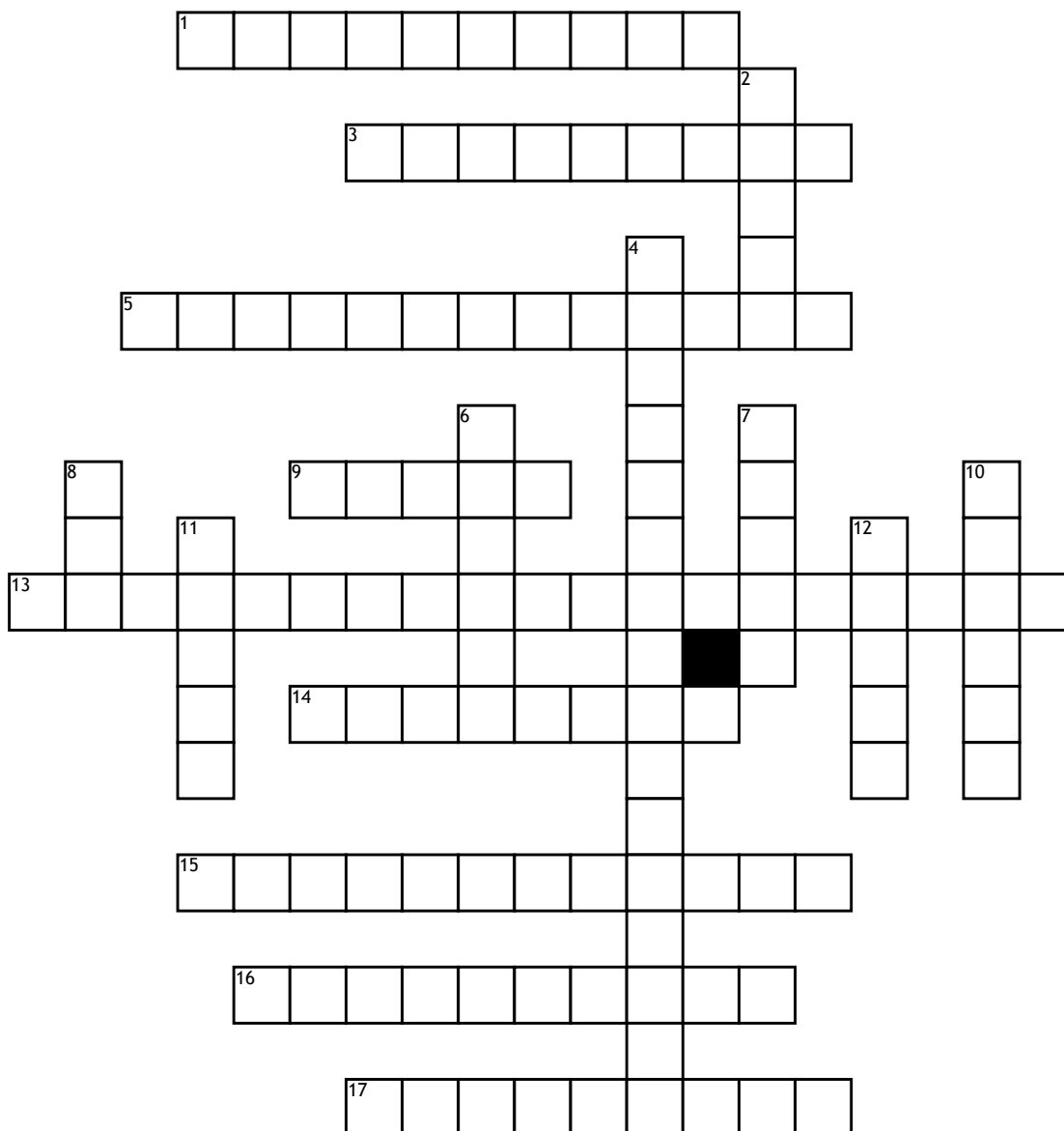


Group 0



Across

1. The of the noble gases decrease as their atomic numbers increase
3. means that the gases exist as single, unbonded atoms
5. The increase as you go down the group
9. What form are all Group 0 elements in at room temperature?
13. Group 0 atoms do not (related to electron structure)

14. Group 0 contains elements

15. They are

16. Another name for Group 0

17. What is the atomic of Krypton?

Down

2. Most elements in Group 0 have electrons in their outer shell

4. Who discovered/isolated most of the Group 0 elements?

6. Which element has a different number of electrons to the others?

7. It's very hard to get them to

8. Helium has electrons in its outer shell

10. They exist as monatomic gases - single atoms not to each other

11. has the second highest boiling point

12. Ar is the symbol for