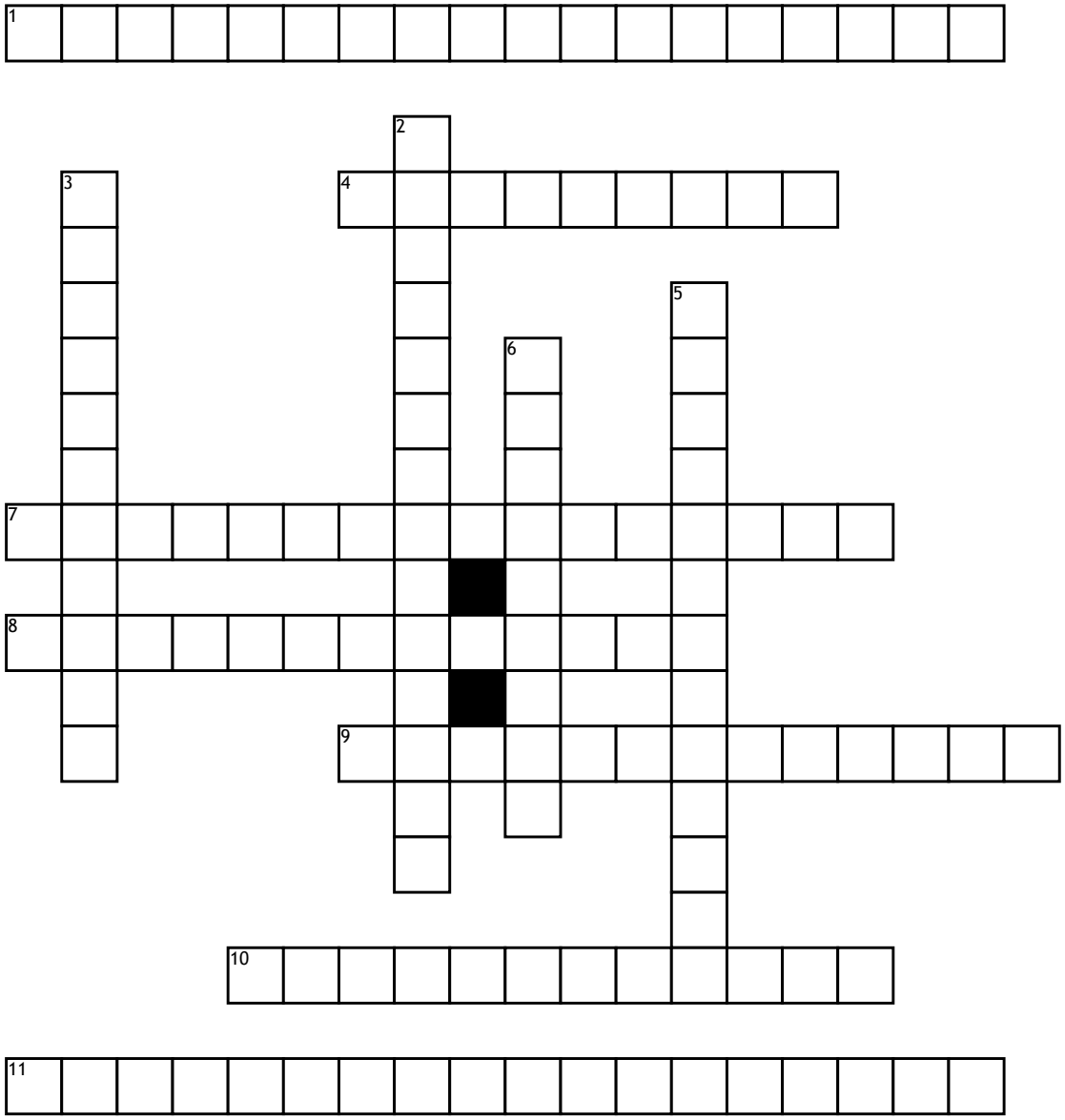
Name: _____ Date: _____

Electrode Potentials



Across

1. The e.m.f of a half-cell compared with a standard hydrogen half cell at 298K, 1 moldm-3 and 100kPa

4. Loss of electrons

7. Type of fuel cell

8. A type of reaction where Reduction and Oxidation occurs

9. Rechargeable cell

10. Equations representing either reduction or oxidation

11. 100 kPa, 298K, 1moldm-3

Down

2. A reagent that oxidises another species

3. Non-rechargeable cell

5. A reagent that reduces another species

6. Gain of electrons