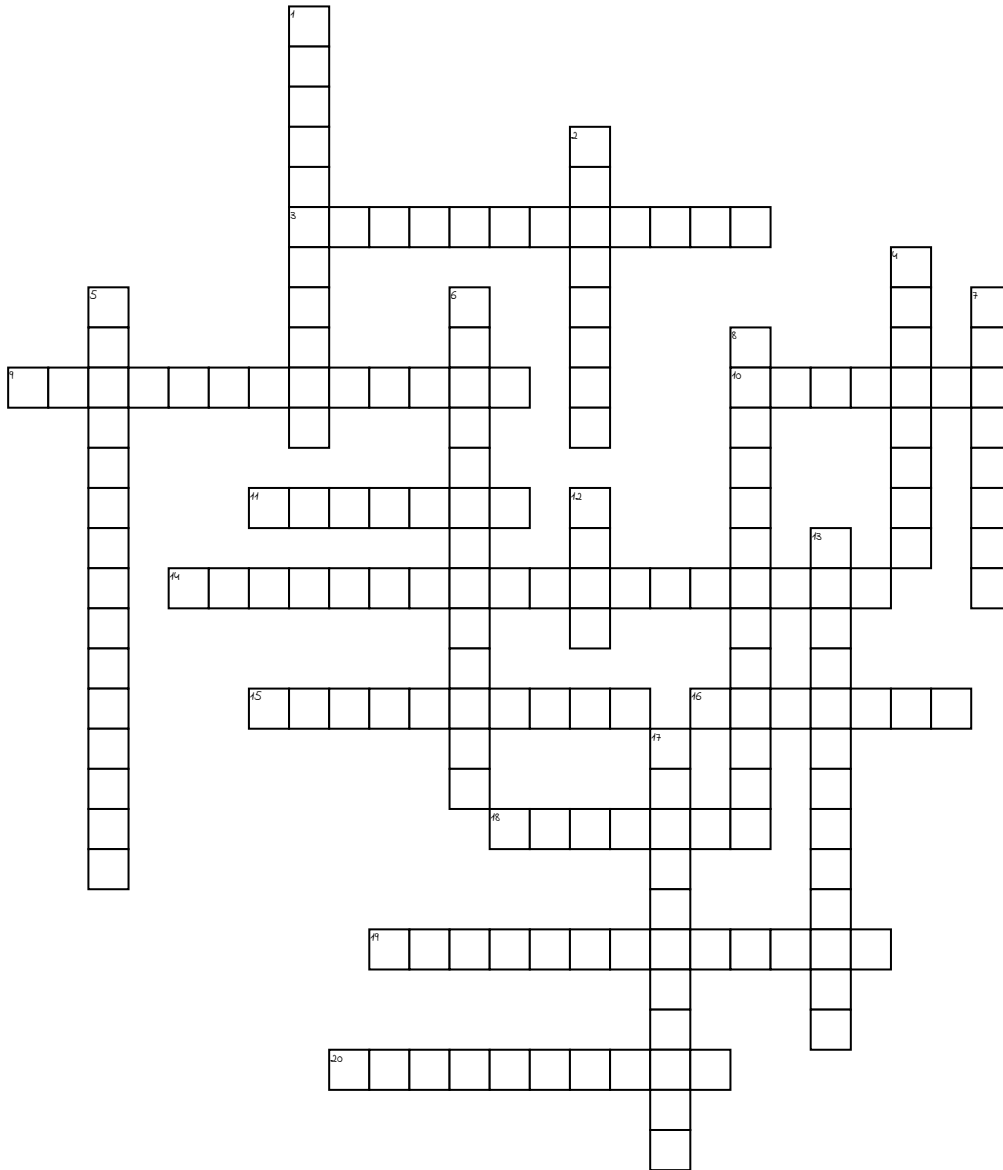


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# ♥ Chemistry vocab ♥



## Across

**3.** the number of protons in the nucleus of an atom, which determines the chemical properties of an element and its place in the periodic table.

**9.** any of the element's lithium, sodium, potassium, rubidium, cesium, and francium, occupying group 1A (1) of the periodic table. They are very reactive, electropositive, monovalent metals forming strongly alkaline hydroxides.

**10.** is a column of elements in the periodic table of the chemical elements.

**11.** is an element that readily forms positive ions (cations) and has metallic bonds

**14.** any of the set of metallic elements occupying a central block (Groups IVB-VIII, IB, and IIB, or 4-12) in the periodic table, e.g. iron, manganese, chromium and copper. Chemically they show variable valence and a strong tendency to form coordination compounds, and many of their compounds are colored.

**15.** a stable subatomic particle with a charge of negative electricity, found in all atoms and acting as the primary carrier of electricity in solids.

**16.** a subatomic particle of about the same mass as a proton but without an electric charge, present in all atomic nuclei except those of ordinary hydrogen.

**18.** a stable subatomic particle occurring in all atomic nuclei, with a positive electric charge equal in magnitude to that of an electron, but of opposite sign.

**19.** which is created by an electron gain and is attracted to the anode.

**20.** is a chemical element that mostly lacks the characteristics of a metal.

## Down

**1.** an element (e.g. arsenic, antimony, or tin) whose properties are intermediate between those of metals and solid nonmetals or semiconductors.

**2.** the central and most important part of an object, movement, or group, forming the basis for its activity and growth.

**4.** each of two or more forms of the same element that contain equal numbers of protons but different numbers of neutrons in their nuclei, and hence differ in relative atomic mass but not in chemical properties; in particular, a radioactive form of an element.

**5.** any of the divalent strongly basic metals of group II of the periodic table comprising beryllium, magnesium, calcium, strontium, barium, and radium.

**6.** any of the element's lithium, sodium, potassium, rubidium, cesium, and francium, occupying Group 1A (1) of the periodic table. They are very reactive, electropositive, monovalent metals forming strongly alkaline hydroxides.

**7.** is the name given to a horizontal row of the periodic table.

**8.** which is created by electron loss and is attracted to the cathode in electrolysis, or as an anion

**12.** an atom or molecule with a net electric charge due to the loss or gain of one or more electrons

**13.** is the region of negative charge surrounding an atomic nucleus that is associated with an atomic orbital.

**17.** a unit of mass used to express atomic and molecular weights, equal to one-twelfth of the mass of an atom of carbon-12. It is equal to approximately  $1.66 \times 10^{-27}$  kg.