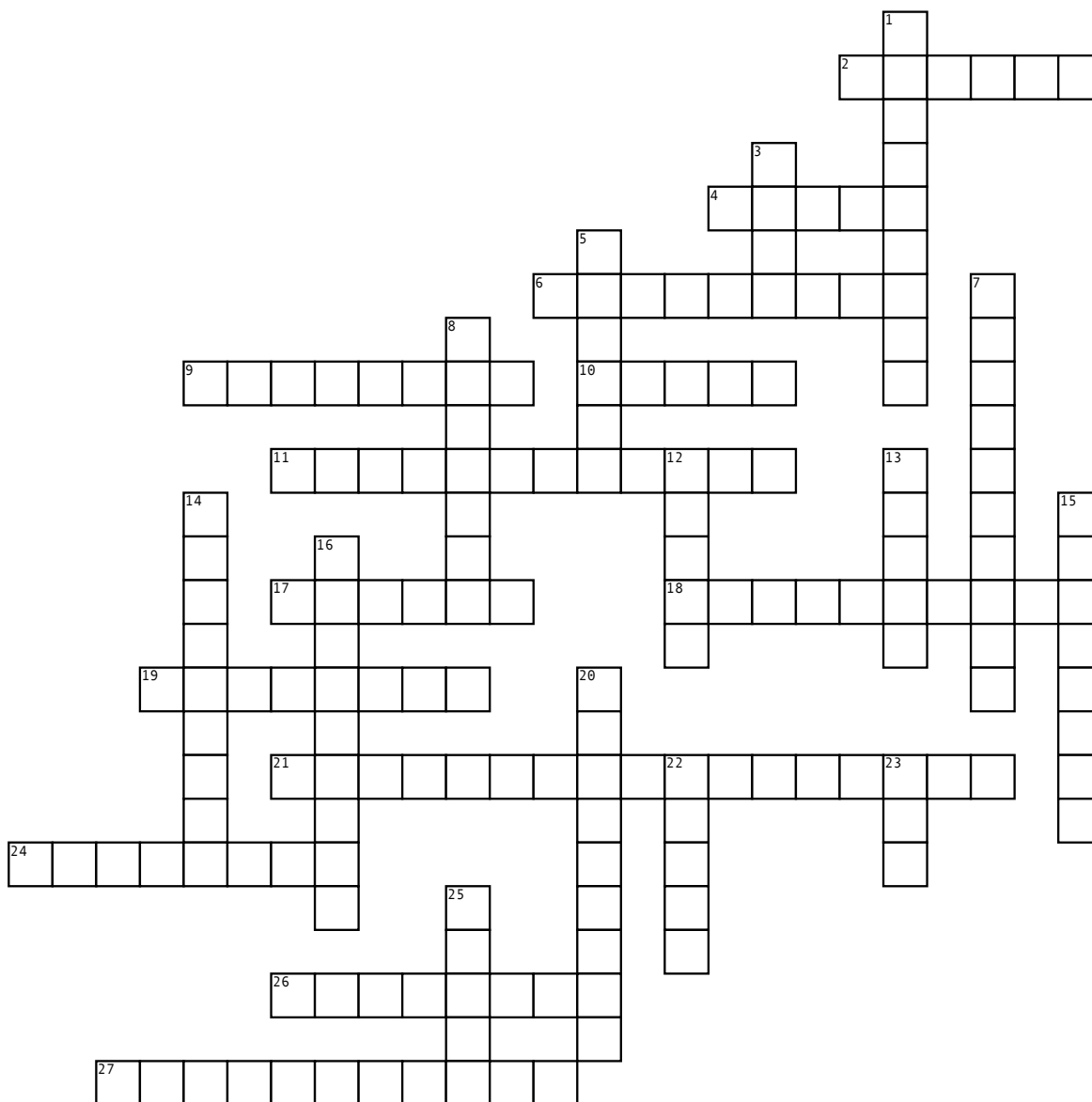


Name: _____

Date: _____

Chemistry Units 6&7



Across

2. Horizontal row on the periodic table.

4. Elements wish to fulfill the octet rule in order to become ___ gases.

6. These are gained or lost in atoms.

9. ___ bipyramidal shape have 6 atoms in the molecule.

10. Potassium is isoelectronic with what noble gas?

11. The ability for metals to become different shapes.

17. The elements in the periodic table are arranged in order of increasing atomic ___.

18. ___ energy is the minimum amount of energy required to create a stable atom.

19. Ionic compounds are made up of ____, which are very large structures consisting of alternating + and - ions.

21. How likely an atom is to attract atoms to itself.

24. A chemical bond that involves the sharing of electron pairs between two nonmetallic atoms.

26. Electronegativity will ___ as you move across a period.

27. The shape of molecules with five atoms.

Down

1. As you move across a period (left to right), the size of an atom ___.

3. A double covalent bond has this many shared electrons.

5. Group 1 on the periodic table is also known as the ___ metals.

7. The elements in the d block of the periodic table are also known as this

8. Electrons in the outermost shell.

12. Bond that involves a transfer of electrons. One gains and the other loses.

13. Electrons shared unequally in a covalent bond.

14. The chemical formula for an ionic compound is arranged in the smallest whole-number ratio is known as the ___ formula.

15. Electrons shared equally in a covalent bond.

16. The ability of an atom to become wire.

20. First scientist to rearrange the periodic table

22. Vertical column on the periodic table.

23. An atom with a positive or negative charge.

25. Valence Shell Electron Pair Repulsion Theory: The electron pairs (both shared and unshared) in the outermost energy level try to get as far apart from each other as possible. Determines shape.