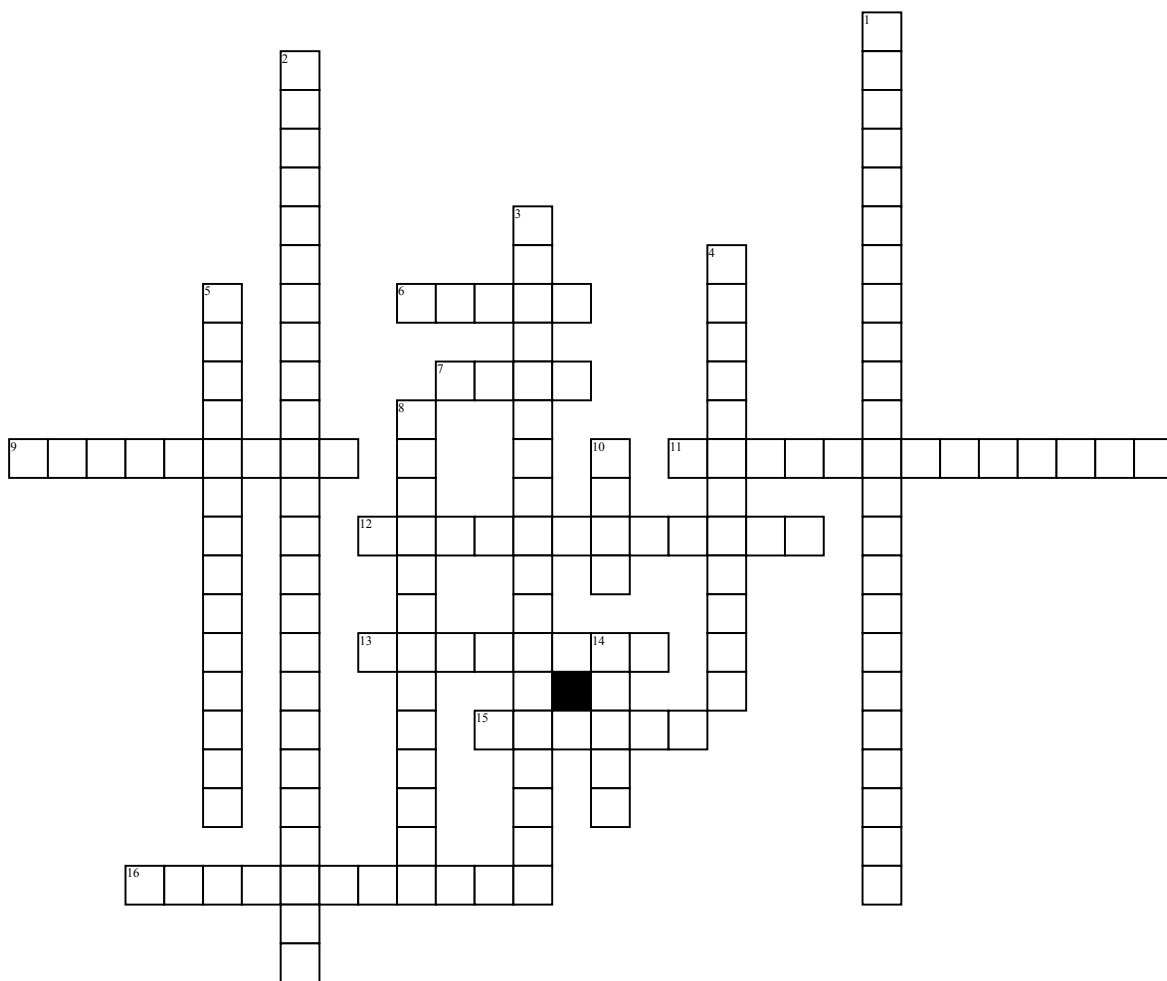


# Chemistry Unit 6



## Across

**6.** a solid material that is typically hard, shiny, malleable, fusible, and ductile, with good electrical and thermal conductivity (e.g., iron, gold, silver, copper, and aluminum, and alloys such as brass and steel).

**7.** the lowest part or edge of something, especially the part on which it rests or is supported

**9.** (of a letter, figure, or symbol) written or printed below the line

**11.** An ion composed of two or more atoms  
**12.** 1 : the monovalent anion  $\text{OH}^-$  consisting of one atom of hydrogen and one of oxygen. 2 : an ionic compound with an element or group

**13.** Is a chemical element that mostly lacks metallic attributes.

**15.** A positive ion. An atom that has lost an electron

**16.** s recommended by IUPAC as a general term for all ions of hydrogen and its isotopes. Depending on the charge of the ion, two different classes can be distinguished: positively charged ions and negatively charged ions

## Down

**1.** sometimes called Proust's law or the law of definite composition, or law of constant composition states that a given chemical compound always contains its component elements in fixed ratio (by mass) and does not depend on its source and method of preparation

**2.** is one of the basic laws of stoichiometry used to establish the atomic theory, alongside the law of conservation of mass (matter)

**3.** An ion composed of a transition metal. Has lost a number of electrons.

**4.** An ion made of one element

**5.** Is made from exactly two different elements.

**8.** also known as a molecular ion, is a charged chemical species (ion) composed of two or more atoms covalently bonded or of a metal complex that can be considered to be acting as a single unit.

**10.** a chemical substance that neutralizes alkalis, dissolves some metals, and turns litmus red; typically, a corrosive or sour-tasting liquid of this kind. Often contrasted with alkali and base.

**14.** A negative ion. An atom that has gained an electron