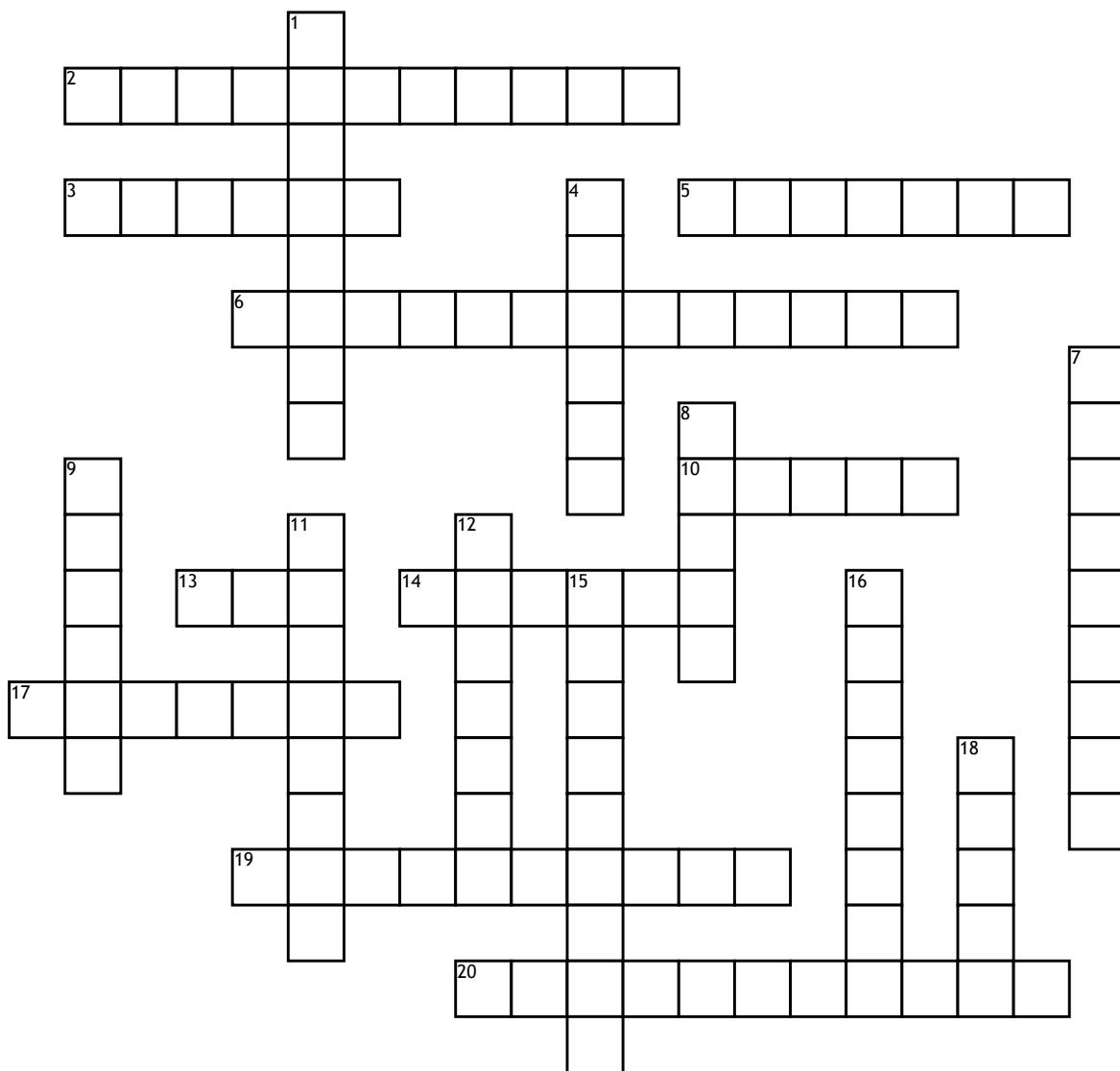


Name: _____

Date: _____

Chemistry Explained



Across

2. Increasing this can cause the rate of reaction to increase

3. The name given to a positive ion

5. For a reaction to occur particles must _____ into each other with enough energy

6. Hydrogen peroxide -> Water + Hydrogen

10. The type of ion that all non-metals form.

13. The number of electrons that Calcium is able to lose

14. The name of the ion that iodine forms

17. Protons and neutrons make up this part of the atom

19. The elements in Group VIII of the periodic table are sometimes called...

20. An insoluble solid formed when two solutions were mixed together.

Down

1. The most reactive group of non-metals

4. What is the name of the positively charged particle in the nucleus of an atom

7. Ionic compounds are combinations of metals and ...

8. An aqueous solution is mostly made up of this.

9. The element that has an electron configuration of 2.8.6

11. A liquid mixture in which an ionic compound has been dissolved.

12. Nitrates are always this.

15. For metals, reactivity _____ as you move down a group.

16. Is not used up by the reaction but can speed it up

18. The number of electrons that can exist in the second shell