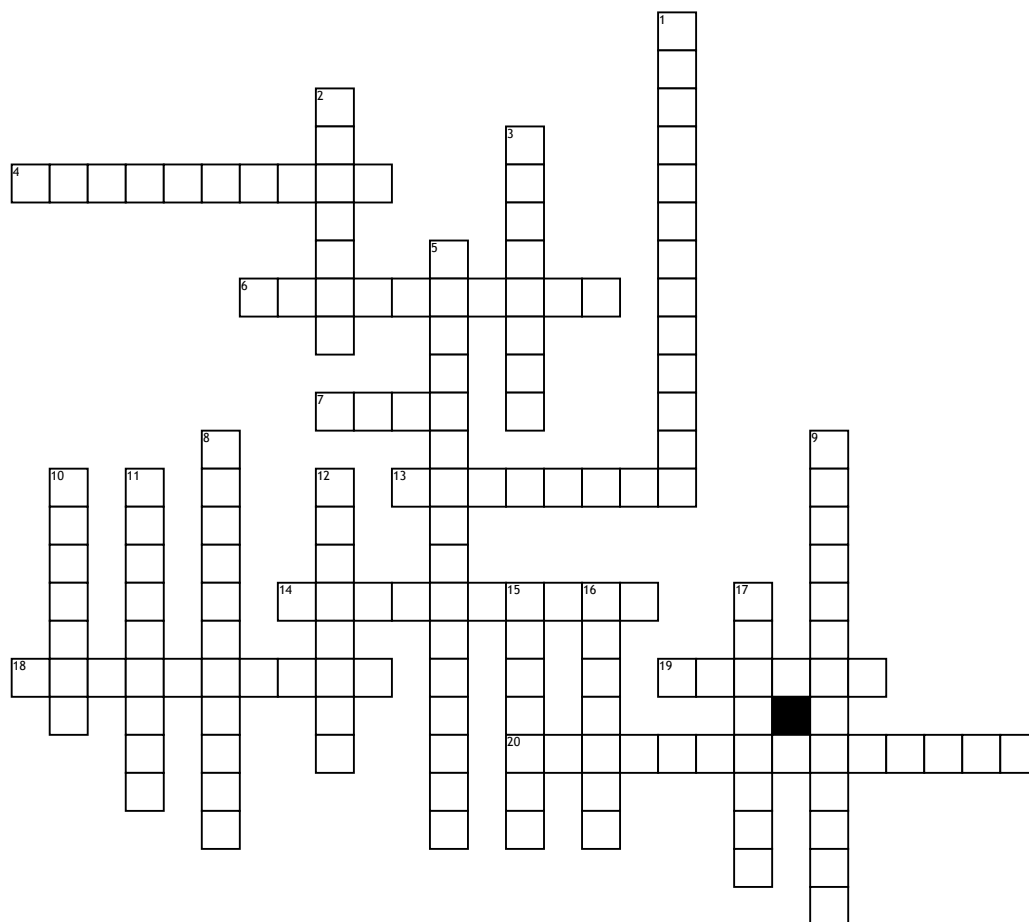


Chemical Reactions



Across

4. not forming a homogeneous mixture when added together.

6. the ability to be dissolved, especially in water.

7. any chemical compound formed from the reaction of an acid with a base, with all or part of the hydrogen of the acid replaced by a metal or other cation.

13. a property of a thermodynamic system, is equal to the system's internal energy plus the product of its pressure and volume.

14. a heterogeneous mixture that contains solid particles sufficiently large for sedimentation

18. certain molecular compounds in which hydrogen is combined with a second nonmetallic element.

19. the minor component in a solution, dissolved in the solvent.

20. It is the remain after a chemical reaction has reached equilibrium.

Down

1. the relationship between the relative quantities of substances taking part in a reaction or forming a compound, typically a ratio of whole integers.

2. able to be dissolved, especially in water.

3. the property of two substances to mix in all proportions, forming a homogeneous solution.

5. the substance that is totally consumed when the chemical reaction is completed.

8. the amount of product obtained in a chemical reaction.

9. the abundance of a constituent divided by the total volume of a mixture.

10. mixture in which one substance of microscopically dispersed insoluble or soluble particles is suspended throughout another substance

11. the ratio between the amounts in moles of any two compounds involved in a chemical reaction

12. a liquid mixture in which the minor component (the solute) is uniformly distributed within the major component (the solvent).

15. able to dissolve other substances.

16. an inorganic acid whose molecules contain oxygen, such as sulfuric or nitric acid.

17. used to express the concentration of a solution. Also known as molar concentration.

Word Bank

oxyacid

solubility

solution

immiscible

actual yeild

excess reactant

soluble

binary acid

miscible

enthalpy

salt

solvent

colloid

mole ratio

solute

suspension

limiting reactant

stoichiometry

concentration

molarity