Name: _____ Date: _____

Chapter 6, 7, & 8 Vocabulary Matching

1. a mutual attraction between two atoms resulting from a redistribution of their outer electrons

2. the complete transfer of valence electron/electrons between atoms

3. a chemical bond that involves the sharing of electron pairs between atoms

4. a type of bond that occurs when two atoms share a pair of electrons with each other and the shared electrons glue two or more atoms together to form a molecule

5. a chemical bond in which the electrons required to form a bond is unequally shared between two atoms

6. a covalent bond between two atoms where the electrons forming the bond are unequally distributed which causes the molecule to have a slight electrical dipole, where one end is slightly positive and the other is slightly negative

7. a group of atoms bonded together, representing the smallest fundamental unit of a chemical compound that can take part in a chemical reaction

8. also known as a covalent compound where the atoms share electrons through covalent bonds

9. tells us the number of atoms of each element in a compound and contains the symbols of the atoms of the elements present in the compound as well as how many there are for each element in the form of subscripts

10. a formula giving the number of atoms of each of the elements present in one molecule of a specific compound

11. the amount of energy required to break apart a mole of molecules into its component atoms

12. only valence electron of an atom of particular elements are shown, indicated by dots placed around the elements' symbol

13. diagrams that show the bonding between atoms of a molecule, and the lone pairs of electrons that may exist in the molecule

A. Multiple bond

B. Structural formula

C. Electron-dot notation

D. Lattice energy

E. Resonance

F. Ionic compound

G. Polyatomic ion

H. Lewis structure

I. Molecular Compound

J. Polar Covalent bond

K. Malleability

L. Molecule

M. Ionic Bonding

14. a formula which shows the arrangement of atoms in the molecule of a compound

15. a chemical bond between two atoms involving two valence electrons in which the atoms share one pair of electrons where the bond forms

16. where two or more electron pairs are shared between two atoms

17. describing delocalized electrons within certain molecules or polyatomic ions where the bonding cannot be expressed by one single Lewis structure

18. a chemical compound composed of ions held together by electrostatic forces

19. the empirical formula of any ionic or covalent network solid compound used as an independent entity for stoichiometric calculations

20. the energy required to separate a mole of an ionic solid into gaseous ions

21. a charged chemical species (ion) composed of two or more atoms covalently bonded or of a metal complex that can be considered to be acting as a single unit

22. type of chemical bond formed between positively charged atoms in which the free electrons are shared among a lattice of cations

23. the ability of a substance, usually a metal, to be deformed or molded into a different shape

24. the physical property of a material associated with the ability to be hammered thin or stretched into wire without breaking

25. a model used to predict the geometry of molecules based on minimizing the electrostatic repulsion of a molecule's valence electrons around a central atom

N. Single bond

O. Chemical formula

P. Ductility

Q. Bond energy

R. Chemical Bond

S. Formula unit

T. VSPER theory

U. Covalent Bonding

V. Polar

W. Molecular formula

X. Metallic bonding

Y. Nonpolar-covalent bond