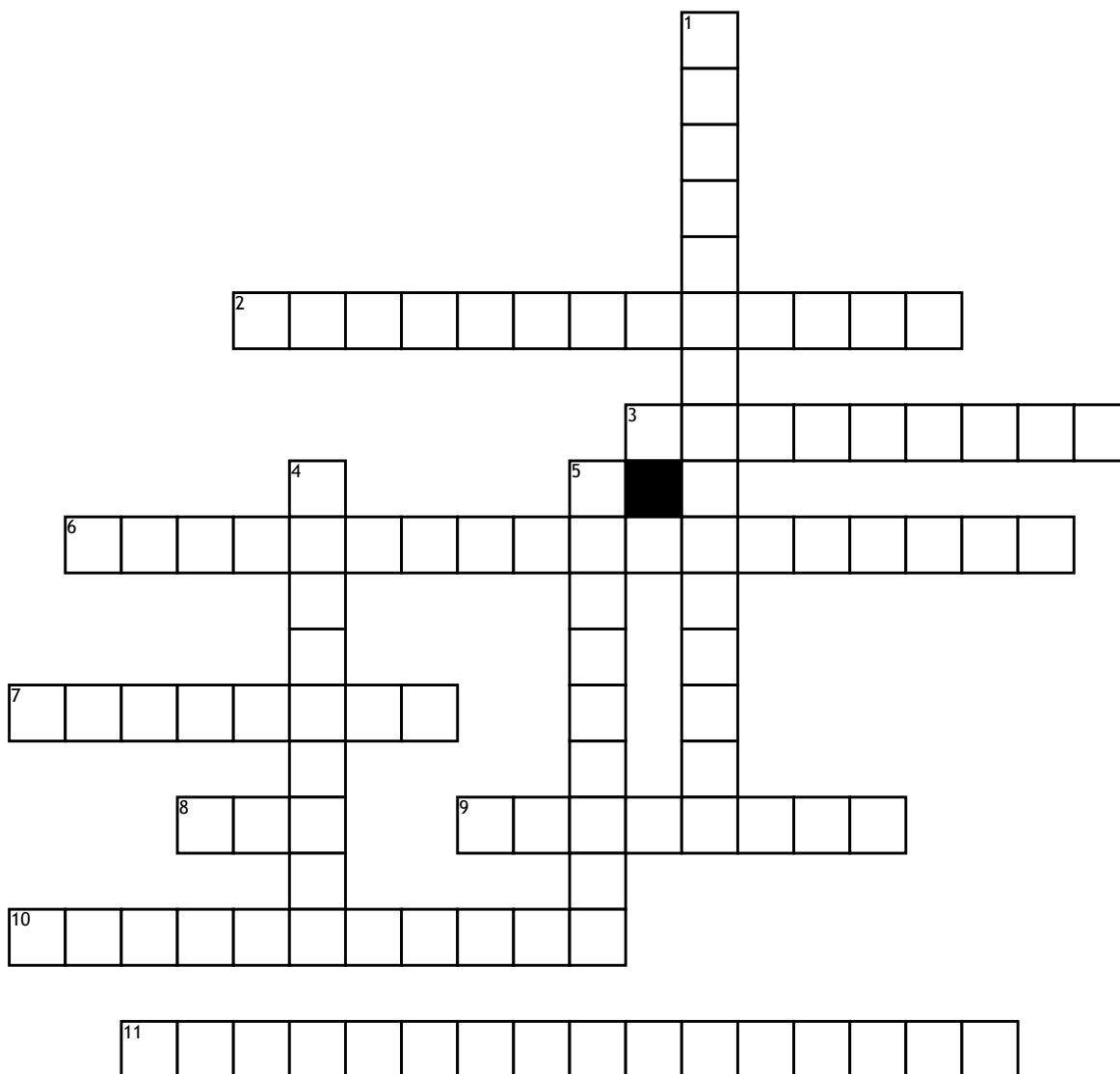


# Chapter 17 chemical equilibrium



## Across

2. applies to reactants in which reactants and products are in different phases
3. an ion that is common to two or more ionic compounds in solution
6. when the equation for a reversible reaction is written in the opposite direction, the equilibrium constant becomes the reciprocal of the original equilibrium constant
7. if  $K > 1$ , it will favor the-
8. A trial value of solubility product that can be compared with  $K_{sp}$  is the
9. the equilibrium constant expression for the dissolving of a sparingly soluble compound is the solubility product-

10. applies to reactants in which all reacting species are in the same phase
11. states that at a given temperature, the balance between reactants and products is a constant value, for a particular reaction

## Down

1. the shift in equilibrium caused by the addition of a compound having an ion in common with the dissolved substance. □□□
4. if stress is applied to a system at equilibrium, the system shifts in the direction that relieves the stress is called L Chatlier's-
5. if  $K < 1$ , it will favor the-