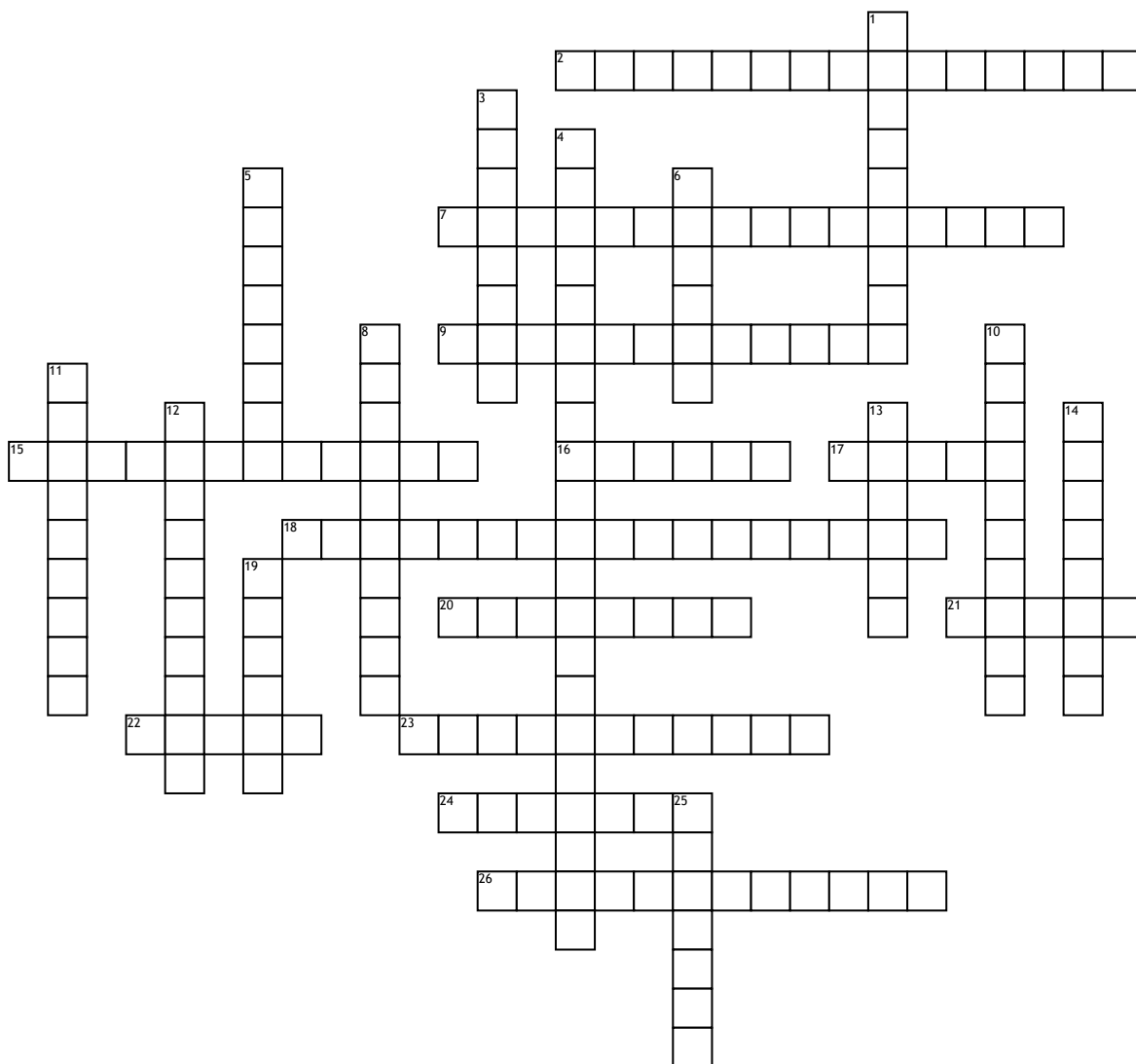


Name: _____ Date: _____ Period: _____

CHEMISTRY REVIEW CROSSWORD



Across

- The electron configuration with arrows representing electrons
- The group number can also tell you the number of _____.
- Are the same as the number of each period, the rings around the nucleus are also called _____.
- Is the same number as the element's protons and neutrons
- A positively charged ion, "meow"
- each column on the periodic table; from top to bottom
- the ability to attract electrons AWAY from other atoms; the "hottest" elements on the periodic table
- An element's symbol surrounded by dots

- Type of bond formed between a metal and nonmetal
- A Negative ION
- The explanation as to why lone pairs and bonded atoms want to stay as far apart as possible from each other.
- Subatomic particle with a NEUTRAL charge
- The distance between the nucleus and the outside ring of an atom

Down

- When they bond together, they like to SHARE their electrons
- Type of bond formed between a nonmetal and a nonmetal
- 1s² is the _____ for Helium
- Subatomic particle with a NEGATIVE charge

- When these bond with another element, they positively lose their electrons
- 4 electrons being shared
- 6 electrons being shared
- The rule of 8 electrons
- 2 electrons being shared
- Subatomic particle with a POSITIVE charge
- Chlorine ion's "married" name
- Each row on the periodic table, from left to right
- Made up of protons and neutrons