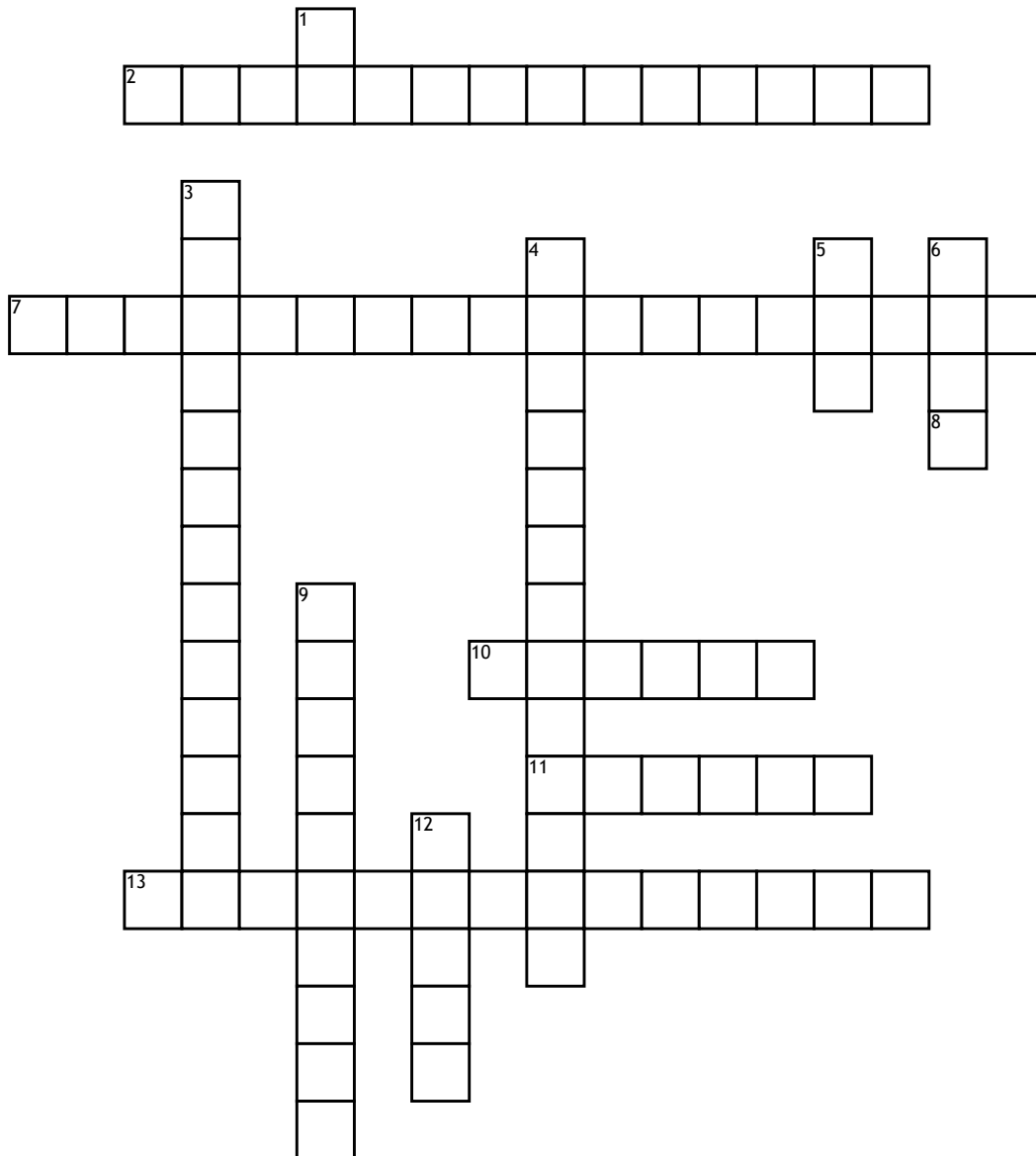


# Acids and Bases



**Across**

- 2. states that an acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solutions
- 7. states that an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor
- 8. both acid and base ionization constant are represented by the letter

10. \_\_\_\_ acids completely ionize

- 11. solution that contains more hydrogen ions than hydroxide ions
- 13. \_\_\_\_\_ reaction will produce salt and water when an acid and a base are mixed

**Down**

- 1. -log[H+] is used to find
- 3. when an acid donates a hydrogen ion to a base is called

4. base accepts a hydrogen ion from an acid is called

- 5. -log[OH-] is used to find
- 6. \_\_\_\_ acids only partially ionize
- 9. when substances acts as both acids and bases is called
- 12. solution that contains more hydroxide ions than hydrogen ions