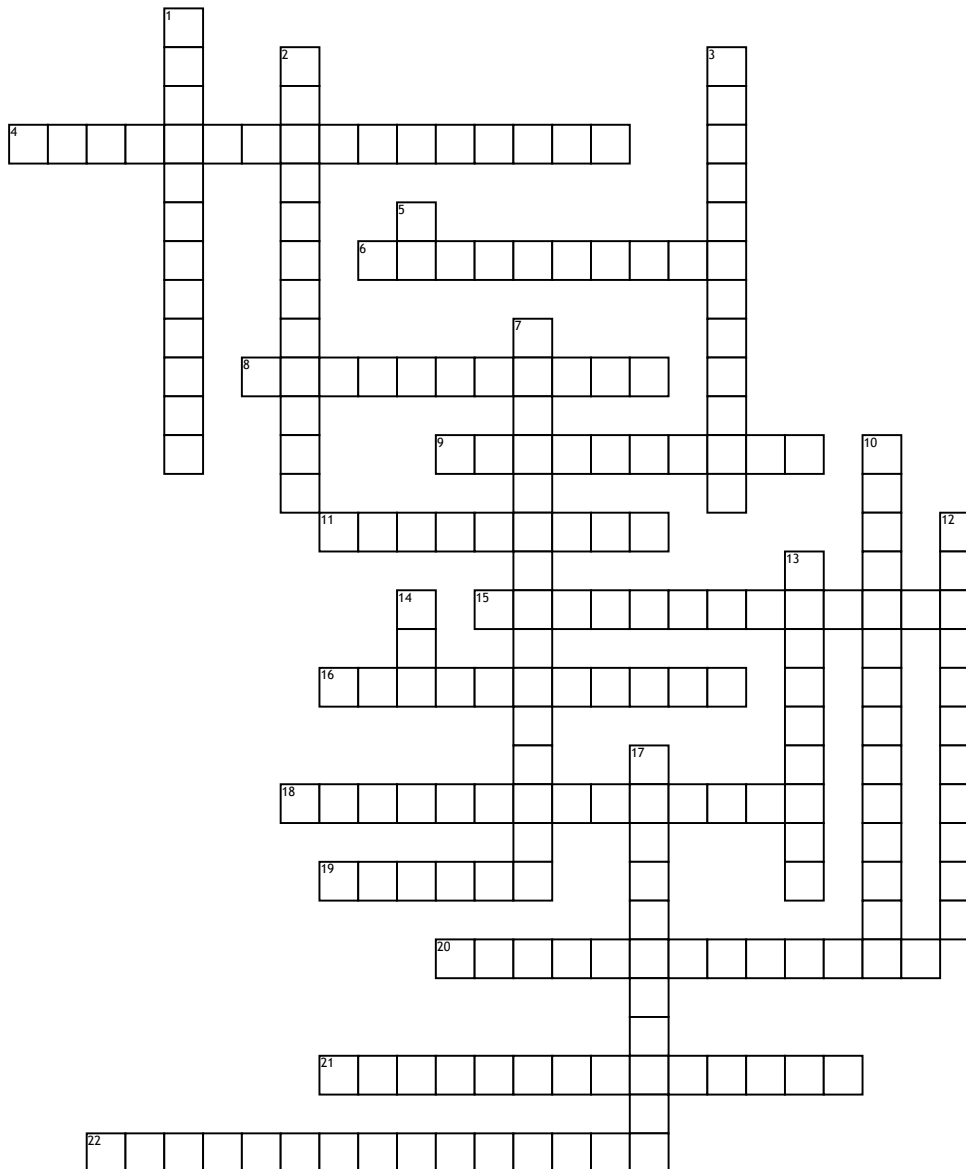


Acids & Bases



Across

4. Have $\text{pH} = 7$
 6. $\text{pH} = -\log[\text{H}^+]$
 8. Acids that ionize completely in solution.
 9. Chemicals that change color in the presence of acids or bases.
 11. Acids that only ionize partially in solution.
 15. The species produced when a base accepts a hydrogen ion to form an acid.
 16. $\text{pOH} = -\log[\text{OH}^-]$
 18. Low pOH and high pH
 19. An indicator that is used to determine if a solution is acidic or basic. Red litmus turns blue for bases, while blue litmus turns red for acids.

20. The species produced when an acid donates a hydrogen ion to form a base.

21. Acid contains H and dissociates to produce H^+ ions in aqueous solution, while a base contains OH and dissociates to produce OH^- ions in aqueous solution.

22. Low pH and high pOH

Down

1. OH^-
 2. When acids and bases ionize - fall apart - in solution to form electrolyte solutions.
 3. H_3O^+ (can be used interchangeably with H^+)

5. A measure of the strength of an acid or base solution which is based on the amount of H^+ ion.

7. Have $\text{pH} < 7$

10. Have $\text{pH} > 7$

12. H^+

13. Bases that ionize only partially in dilute aqueous solution to form the conjugate acid and hydroxide ions.

14. A measure of the strength of an acid or base solution which is based on the amount of OH^- ion.

17. Bases that dissociate entirely into metal ions and hydroxide (OH^-) ions in aqueous solution (Arrhenius base).