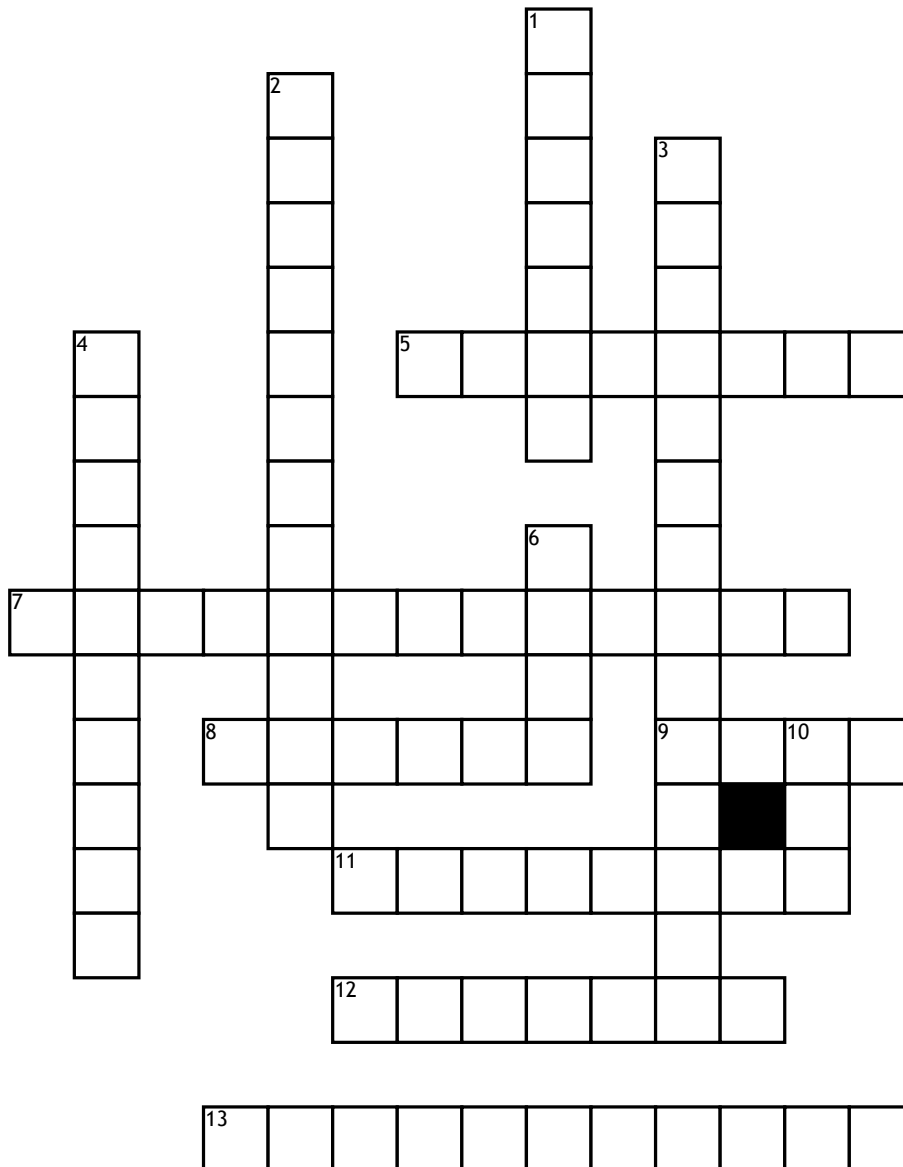


# Acids & Bases Key Terms



## Across

5. When a less concentrated solution is prepared by adding additional solvent.
7. The ratio of the amount of solute to the volume of solvent
8. The substance that changes state when a solution is prepared or is present in the smallest amount.
9. A chemical compound that disassociates into a hydrogen ion(s) and a negatively charged ion when mixed with water
11. A mixture of two or more substances that is homogenous or the same throughout.

12. The substance that does not change when a solution is prepared or is present in the largest amount. In the context these investigation, this is usually water.

13. When an acid is mixed with water, it disassociated into a negatively charged ion and this

## Down

1. A linear scales from 0-14 that is a measure of the hydrogen ion concentration

2. When a base is mixed with water, it disassociated into a positively charged ion and this

3. When an acid and a base react to form a water molecule(s) and reduce the hydrogen ion concentration

4. When a chemical compound is dissolved in a solvent, usually water, and separated into two or more ions.

6. A chemical compound that disassociated into hydroxide ion(s) and a positively charged ion when mixed with water.

10. A chemical compound that has a negative or positive charge. A cation is positively charged and an anion is negatively charged.