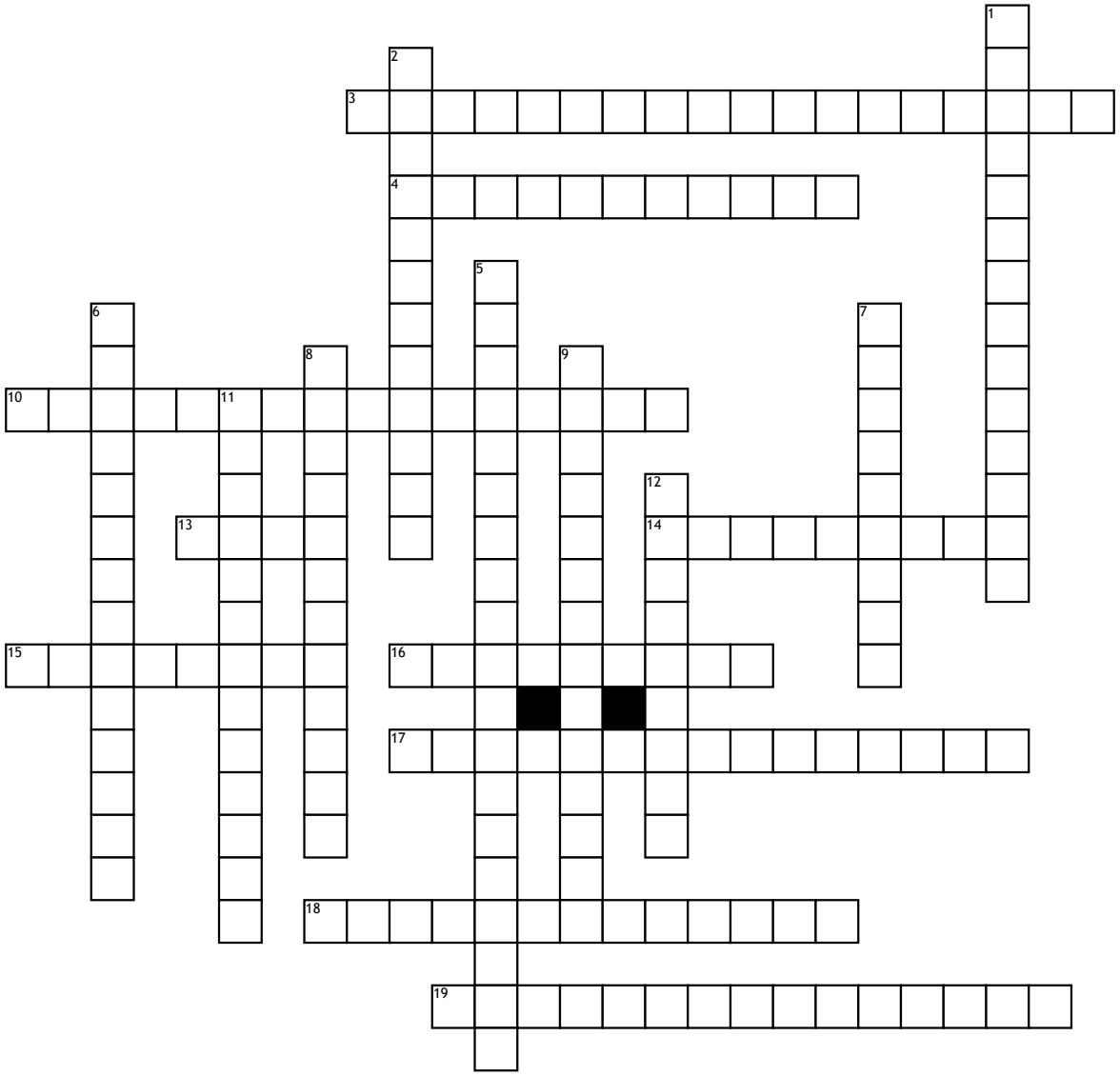
Name: _____ Date: _____

AP Chem Crossword



Across

3. a technique used to control the corrosion of a metal surface by making it the cathode of an electrochemical cell

4. law stating that the amount of any substance deposited or liberated during electrolysis is proportional to the quantity of electric charge passed and to the equivalent weight of the substance

10. a cell consisting of an electrolyte, its container, and two electrodes, in which the electrochemical reaction between the electrodes and the electrolyte produces an electric current

13. the SI unit of electromotive force, the difference of potential that would drive one ampere of current against one ohm resistance

14. the process or result of oxidizing or being oxidized

15. one of the two electrodes in a galvanic cell or simple battery

16. a conductor through which electricity enters or leaves an object, substance, or region

17. describes the degree of oxidation of an atom in a chemical compound

18. a substance that tends to bring about reduction by being oxidized and losing electrons.

19. the process of using electrolysis to increase the purity of a metal extracted from its ore

Down

1. a substance that tends to bring about oxidation by being reduced and gaining electrons

2. is either the oxidation or reduction reaction component of a redox reaction

5. a device capable of either generating electrical energy from chemical reactions or using electrical energy to cause chemical reactions

6. a process that uses an electric current to reduce dissolved metal cations so that they form a thin coherent metal coating on an electrode

7. the process or result of reducing or being reduced.

8. a technique that uses a direct electric current to drive an otherwise non-spontaneous chemical reaction

9. a type of titration based on a redox reaction between the analyte and titrant

11. any chemical reaction in which the oxidation number of a molecule, atom, or ion changes by gaining or losing an electron

12. a natural process, which converts a refined metal to a more chemically-stable form, such as its oxide, hydroxide, or sulfide